

## Chemistry – Unit 5 Worksheet 3

### Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

1. Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

$$32.0 \text{ g} \times \frac{1 \text{ mole}}{79.9 \text{ g}} = 0.401 \text{ mole Br}$$

1:2 ratio

$$4.9 \text{ g} \times \frac{1 \text{ mole}}{24.31 \text{ g}} = 0.20 \text{ mole Mg}$$

 $\text{MgBr}_2$ 

2. What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?

$$95.2 \text{ g} - 40.8 \text{ g} = 54.4 \text{ g O}$$

$$40.8 \text{ g} \times \frac{1 \text{ mole}}{12.0 \text{ g}} = 3.40 \text{ moles C}$$

1:1 ratio

$$54.4 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 3.40 \text{ moles O}$$

 $\text{CO}$ 

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

$$9.8 \text{ g} \times \frac{1 \text{ mole}}{14.0 \text{ g}} = 0.70 \text{ moles N}$$

1:1:3 ratio

$$0.70 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.69 \text{ moles H}$$

 $\text{HNO}_3$ 

$$33.6 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 2.10 \text{ moles O}$$

4. A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

$$0.59 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.58 \text{ mole H}$$

1:1 ratio; empirical formula:  $\text{HO}$   
 empirical formula mass: 17.0 g/mol  
 molecular formula mass: 34.0 g/mol  
 mfm is 2x efm so the molecular formula is:  $\text{H}_2\text{O}_2$

$$9.40 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 0.588 \text{ mole O}$$

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