

## Chemistry – Unit 6 Worksheet 1

We have observed evidence that atoms of different elements are different, the most noticeable being in their predictable characteristic chemical behaviour. Atoms of one element cannot be transformed into atoms of another element, which demonstrates that each element's particles are made up of atoms with only one type of nucleus. However, when the elements are split, subatomic particles are revealed as their composition changes. Atoms contain subatomic particles as nuclei. We will now examine the particles that make up the atoms for which there is evidence a nucleus is made up of protons and neutrons, the building blocks of the atom's deep core.

1. Write the formula and draw the particle diagram for each compound. The ratio of ions in each compound is given.

Atomic number of compound	16 electrons 16 protons 1:1 ratio	16 electrons 16 protons 1:1 ratio	16 electrons 17 protons 1:2 ratio	16 electrons 17 protons 1:2 ratio
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particle



Diagram

Atomic number of compound	16 electrons 16 protons 1:1 ratio	16 electrons 16 protons 1:1 ratio	16 electrons 17 protons 1:2 ratio	16 electrons 17 protons 1:2 ratio
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particle



Diagram

Atomic number of compound	16 electrons 16 protons 1:1 ratio	16 electrons 16 protons 1:1 ratio	16 electrons 17 protons 1:2 ratio	16 electrons 17 protons 1:2 ratio
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particle



Diagram

# Chemistry Unit 6 Lesson 1 Answers Key

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