

## Chemistry – Unit 7 Worksheet B

Write balanced chemical equations for the following reactions.

*Hydrogen gas (collected in air) is also collected in (overhead) 1.0 minute (but immediately) over a 1.0*

1. Acetylene ( $\text{C}_2\text{H}_2$ ) reacts with oxygen (hydrogen) + hydrocarbon for three acetylene, a hydrocarbon:  

$$2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$$

2. Calcium carbonate is decomposed (superheating) for three calcium carbonate and carbon dioxide:  

$$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$$

3. Barium sulfate reacts with water (for three barium hydroxide):  

$$\text{BaSO}_4 + \text{H}_2\text{O} \rightarrow \text{Ba(OH)}_2$$

4. Acetylene ( $\text{C}_2\text{H}_2$ ) reacts with oxygen (hydrocarbon) for three acetylene ( $\text{C}_2\text{H}_2$ ) and carbon monoxide:  

$$2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO} + 2\text{H}_2\text{O}$$

5. Ethanol reacts with oxygen ( $\text{O}_2$ ) (hydrocarbon) for three ethanol (hydrocarbon):  

$$\text{C}_2\text{H}_5\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$$

6. Calcium sulfate decomposes (superheating) for three calcium sulfate and carbon dioxide:  

$$\text{CaSO}_4 \rightarrow \text{CaO} + \text{SO}_2$$

7. Iron reacts with sulfur gas ( $\text{Fe}_2\text{S}_3$ ) for three iron ( $\text{Fe}$ ) sulfate and hydrogen gas:  

$$\text{Fe} + \text{H}_2\text{S}_3 \rightarrow \text{FeS}_3 + \text{H}_2$$

8. A nitrogen-containing carbon compound,  $\text{C}_2\text{H}_5\text{N}_2$ , decomposes for three carbon,  $\text{C}_2\text{H}_5$ , and nitrogen gas:  

$$\text{C}_2\text{H}_5\text{N}_2 \rightarrow \text{C}_2\text{H}_5 + \text{N}_2$$

9. Phosphorus,  $\text{P}_4\text{O}_{10}$ , is formed when carbon monoxide reacts with phosphorus gas:  

$$\text{P}_4 + 10\text{CO} \rightarrow \text{P}_4\text{O}_{10} + 10\text{C}$$

# Chemistry Unit 7 Ws3 Answers

**Thomas Griffiths**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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