

7. molar heat of condensation

Part B True-False

8. AT 10. ST 12. NT
9. NT 11. NT

Part C Matching

13. a 15. b 17. d
14. e 16. c

Part D Questions and Problems

18. a. endothermic d. endothermic
b. endothermic e. exothermic
c. exothermic f. exothermic

$$19. \Delta H = 28.3 \frac{\text{g H}_2\text{O}(g)}{1 \text{ mol H}_2\text{O}(g)} \times \frac{1 \text{ mol H}_2\text{O}(g)}{18.0 \text{ g H}_2\text{O}(g)} \\ \times \frac{6.01 \text{ kJ}}{1 \text{ mol H}_2\text{O}(g)} = 9.45 \text{ kJ}$$

$$20. \Delta H = 5.53 \frac{\text{mol NH}_4\text{NO}_3(g)}{1 \text{ mol NH}_4\text{NO}_3(g)} \\ \times \frac{25.7 \text{ kJ}}{1 \text{ mol NH}_4\text{NO}_3(g)} = 142 \text{ kJ}$$

Section 17.4

Part A Completion

- sum
- enthalpy
- indirectly
- changed (reversed)
- standard heat of formation
- change
- one
- ΔH_f°
- zero
- subtracting

Part B True-False

11. NT 13. AT 15. NT
12. NT 14. NT

Part C Matching

16. b 18. a 20. c
17. c 19. d

Part D Questions and Problems



$$22. \Delta H_f^\circ (\text{products}) \\ = 3 \frac{\text{mol CO}_2(g)}{1 \text{ mol CO}_2(g)} \times \frac{-393.5 \text{ kJ}}{1 \text{ mol CO}_2(g)} \\ = -1181 \text{ kJ} \\ \Delta H_f^\circ (\text{reactants}) \\ = 3 \frac{\text{mol CO}(g)}{1 \text{ mol CO}(g)} \times \frac{-110.5 \text{ kJ}}{1 \text{ mol CO}(g)} + (-822.1 \text{ kJ}) \\ = -1154 \text{ kJ} \\ \Delta H^\circ = \Delta H_f^\circ (\text{products}) - \Delta H_f^\circ (\text{reactants}) \\ = -1181 \text{ kJ} - (-1154 \text{ kJ}) \\ = -27 \text{ kJ}$$

Practice Problems

Section 17.1

- $200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}} \\ = 836.8 \text{ kJ}$
- $C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0^\circ\text{C}} = 1.40 \text{ cal/(g} \times ^\circ\text{C)}$
- $\Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J/(g} \times ^\circ\text{C)}} = 6.0^\circ\text{C}$
- $q = 100.0 \text{ g} \times 120.0^\circ\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \\ = 1.1 \times 10^4 \text{ J}$

Section 17.2

- $\Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \times 10^\circ\text{C} \\ = 6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ}$
- $\Delta H = 15.0 \text{ g Ca(OH)}_2(g) \times \frac{1 \text{ mol Ca(OH)}_2(g)}{74.1 \text{ g Ca(OH)}_2(g)} \\ \times \frac{-65.2 \text{ kJ}}{1 \text{ mol Ca(OH)}_2(g)} = -13.2 \text{ kJ}$
- $\Delta H = 52.4 \text{ g CH}_4(g) \times \frac{1 \text{ mol CH}_4(g)}{16.0 \text{ g CH}_4(g)} \\ \times \frac{-890.2 \text{ kJ}}{1 \text{ mol CH}_4(g)} = 2.93 \times 10^3 \text{ kJ}$
- $4\text{NH}_3 + 5\text{O}_2(g) \rightarrow 4\text{NO}(g) + 6\text{H}_2\text{O}(g) \\ \Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_3(g)} \times 4 \text{ mol NH}_3(g) \\ = -904 \text{ kJ}$

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electron 0.0005486 u

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mass of element mass of isotope 1 in amu mass fraction of isotope 1 mass of isotope 2 mass fraction of isotope 2 bromine has

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protons p n mass element name atomic number number of protons number of neutrons mass number carbon 6 6 6 12 oxygen
8 8 8 16 hydrogen 1 1 0 1 carbon 6 6 8 14 hydrogen 1 1 2 3 nitrogen 7 7

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web the atomic mass of a single atom is simply its total mass and is typically expressed in atomic mass units or amu by definition an atom of carbon with six neutrons carbon 12 has an atomic mass of 12 amu other atoms don't generally have round number atomic masses for reasons that are a little beyond the scope of this article

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