

Section 8.4: Calculations Involving Acidic Solutions

Tutorial 1 Practice, page 513

1. Given: $[\text{HCl}(\text{aq})] = 0.0700 \text{ mol/L}$

Required: $[\text{OH}^-(\text{aq})]$

Analysis: $[\text{H}^+(\text{aq})] = 0.0700 \text{ mol/L}$ because HCl is a strong acid

$$[\text{OH}^-(\text{aq})] = \frac{1.0 \times 10^{-14}}{[\text{H}^+]}$$

$$\text{Solution: } [\text{OH}^-(\text{aq})] = \frac{1.00 \times 10^{-14}}{7.00 \times 10^{-2}}$$

$$[\text{OH}^-(\text{aq})] = 1.43 \times 10^{-13} \text{ mol/L}$$

2. Given: $[\text{HBr}(\text{aq})] = 0.0380 \text{ mol/L}$; $\text{pH} = 2.78$

Required: pH , pOH

Analysis: $[\text{H}^+(\text{aq})] = 0.0350 \text{ mol/L}$ because HBr is a strong acid

$$\text{pH} = -\log[\text{H}^+(\text{aq})]$$

$$\text{pOH} = 14.00 - \text{pH}$$

$$\text{Solution: } \text{pH} = -\log[\text{H}^+(\text{aq})]$$
$$= -\log(3.5 \times 10^{-2}) = 1.46$$

$$\text{pH} = 1.46$$

$$\text{pOH} = 14.00 - 1.46$$

$$= 12.54$$

Tutorial 2 Practice, page 516

1. Given: $[\text{HC}_2\text{H}_3\text{O}_2(\text{aq})] = 0.050 \text{ mol/L}$; $\text{pH} = 2.78$

Required: percentage ionization

Analysis: $[\text{H}^+(\text{aq})] = 10^{-\text{pH}}$

$$\frac{\text{percentage ionization}}{100} = \frac{[\text{H}^+(\text{aq})]}{[\text{HC}_2\text{H}_3\text{O}_2(\text{aq})]}$$

$$\text{Solution: } [\text{H}^+(\text{aq})] = 10^{-2.78}$$

$$[\text{H}^+(\text{aq})] = 0.00166 \text{ mol/L}$$

$$\frac{\text{percentage ionization}}{100} = \frac{0.00166}{0.050}$$

$$\text{percentage ionization} = 3.3 \%$$

2. Given: $[\text{HF}(\text{aq})] = 0.100 \text{ mol/L}$; percentage ionization = 7.8 %

Required: K_a

Analysis: $\text{HF}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{F}^-(\text{aq})$

$$K_a = \frac{[\text{H}^+(\text{aq})][\text{F}^-(\text{aq})]}{[\text{HF}(\text{aq})]}$$

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