

Colligative (collective) properties are vapor pressure lowering, boiling point elevation (ΔT_b), freezing point depression (ΔT_f), and osmotic pressure (Π). **Colligative properties depend only on the number of solute particles and not their chemical identity.** This is best, specifically, the properties of a pure solvent compare with those of a solution:

- The vapor pressure of the solution is lower.
- The boiling point of the solution is higher.
- The freezing (or melting) point of the solution is lower.
- The solution causes osmosis, the flow of solvent from a solution of lower solute concentration to one of higher solute concentration a semipermeable membrane.

Below are the key equations you need for solving problems related to colligative properties.

Key Equations for Colligative Properties

Raoult's Law: Relationship between the Vapor Pressure of a Solution (P_{solution}), the Mole Fraction of the Solvent (χ_{solvent}), and the Vapor Pressure of the Pure Solvent ($P^{\circ}_{\text{solvent}}$)

$$P_{\text{solution}} = \chi_{\text{solvent}} P^{\circ}_{\text{solvent}}$$

$$\chi_{\text{solvent}} = \frac{P_{\text{solution}}}{P^{\circ}_{\text{solvent}}}$$

The Vapor Pressure of a Solution Containing Two Volatile Components

$$P_A = \chi_A P^{\circ}_A \quad P_B = \chi_B P^{\circ}_B$$

$$P_{\text{total}} = P_A + P_B$$

Relationship between Osmotic Pressure (Π), Molarity (M), the Ideal Gas Constant (R), and Temperature (T , in K)

$$\Pi = MRT$$

$$M = \frac{n_{\text{solute}}}{V_{\text{solution}}} \quad L = \text{liters} / \text{mol} \cdot K$$

The dependence of ΔT_f and ΔT_b on the Molarity (m) and Freezing Point Depression Constant (K_f) and Boiling Point Elevation Constant (K_b) respectively:

$$\Delta T_f = m \cdot K_f$$

$$\Delta T_b = m \cdot K_b$$

$$m = \frac{\text{mol (solute)}}{\text{kg (solvent)}}$$

An important trait in colligative properties is the larger the number of the nodes, the stronger the effect. For example, dissolving 2 moles of methanol (nonelectrolyte) will not affect the colligative properties of the solution as much as 2 moles of NaCl (electrolyte) would. This is because NaCl dissociates into ions, so 2 moles of NaCl will produce 4 moles of particles (ions).



Electrolytes have stronger effect on colligative properties of solutions.

Therefore, we divide our summary in two parts: colligative properties of nonelectrolytes solutions and those of electrolyte solutions.

Nonvolatile Nonelectrolyte Solutions

Vapor Pressure Lowering Example: Calculate the vapor pressure of a solution at 25°C that is made by adding 47.5 g of glucose ($C_6H_{12}O_6$) to 340 g of water. The vapor pressure of pure water at 25°C is 23.8 torr.

The vapor pressure of the solution is calculated by the following formula:

$$P_{\text{solution}} = \chi_{\text{solvent}} P^{\circ}_{\text{solvent}}$$

We have the vapor pressure of pure water: $P^{\circ}(\text{H}_2\text{O}) = 23.8 \text{ Torr}$, so the only missing part is the mole fraction of the solvent (water in this case) which is the ratio of the moles of solvent over the total moles of the solvent and the solute.

Calculate the moles and plug the numbers on the equation for the mole fraction of H_2O :

$$\chi_{\text{H}_2\text{O}} = \frac{n(\text{H}_2\text{O})}{n(\text{H}_2\text{O}) + n(\text{C}_6\text{H}_{12}\text{O}_6)} = \frac{15.9 \text{ mol}}{(15.9 + 0.266) \text{ mol}} = 0.983$$

The vapor pressure of the solution would be: $P_{\text{solution}} = 0.983 \times 23.8 \text{ torr} = 23.4 \text{ torr}$

Freezing Point Depression Example: Using the appropriate data in the table, determine the freezing point depression of the solution that contains 24.1 g urea (NH_2CONH_2) in 485 mL of water.

The freezing point depression is calculated by the following formula: $\Delta T_f = m \cdot K_f$

So, we need to first determine the moles of urea, then calculate the molarity and use it in the equation for the ΔT_f .

$$n(\text{NH}_2\text{CONH}_2) = 24.1 \text{ g} \times \frac{1 \text{ mol}}{60.1 \text{ g}} = 0.401 \text{ mol}$$

$$m(\text{H}_2\text{O}) = 485 \text{ mL} \times \frac{1.00 \text{ g}}{1 \text{ mL}} \times \frac{1 \text{ kg}}{1000 \text{ g}} = 0.485 \text{ kg}$$

$$m = \frac{0.401 \text{ mol}}{0.485 \text{ kg}} = 0.827 \text{ m}$$

$$\Delta T_f = K_{fm} = 1.86^\circ\text{C} / \text{m} \times 0.827 \text{ m} = 0.591^\circ\text{C}$$

Boiling-Point and Freezing-Point Constants

Solvent	Freezing point (°C)	K_f (K·kg·mol ⁻¹)	Boiling point (°C)	K_b (K·kg·mol ⁻¹)
acetone	-93.85	2.40	56.3	0.71
benzene	5.5	5.12	80.1	2.53
carbon tetrachloride	-23	29.8	76.5	4.93
cyclohexane	6.5	20.1	80.7	2.79
naphthalene	80.3	6.94	217.7	5.80
phenol	41	7.27	182	3.84
water	0	1.86	100.0	0.51

Chemistry Study Guide Calculations Involving Colligative Properties

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Chemistry Study Guide Calculations Involving Colligative Properties:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

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 - The Rise of Digital Reading Chemistry Study Guide Calculations Involving Colligative Properties
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Calculations Involving Colligative Properties
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Calculations Involving Colligative Properties
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Calculations Involving Colligative Properties
 - Personalized Recommendations
 - Chemistry Study Guide Calculations Involving Colligative Properties User Reviews and Ratings
 - Chemistry Study Guide Calculations Involving Colligative Properties and Bestseller Lists
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 - Chemistry Study Guide Calculations Involving Colligative Properties Public Domain eBooks
 - Chemistry Study Guide Calculations Involving Colligative Properties eBook Subscription Services
 - Chemistry Study Guide Calculations Involving Colligative Properties Budget-Friendly Options
6. Navigating Chemistry Study Guide Calculations Involving Colligative Properties eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Study Guide Calculations Involving Colligative Properties Compatibility with Devices
- Chemistry Study Guide Calculations Involving Colligative Properties Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Calculations Involving Colligative Properties
 - Highlighting and Note-Taking Chemistry Study Guide Calculations Involving Colligative Properties
 - Interactive Elements Chemistry Study Guide Calculations Involving Colligative Properties
- 8. Staying Engaged with Chemistry Study Guide Calculations Involving Colligative Properties
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Calculations Involving Colligative Properties
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Calculations Involving Colligative Properties
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Calculations Involving Colligative Properties
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Calculations Involving Colligative Properties
 - Setting Reading Goals Chemistry Study Guide Calculations Involving Colligative Properties
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Calculations Involving Colligative Properties
 - Fact-Checking eBook Content of Chemistry Study Guide Calculations Involving Colligative Properties
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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