

## Summer Assignment Chapter 2

PROBLEMS pg. 69 #7, 11, 13, 15, 17, 20, 21, 22, 27, 28, 30, 31, 33, 35, 38, 40, 41, 43, 45, 47, 49, 55, 56, 57, 58, 60, 61, 64, 67, 69, 71, 87

- 7 The diameter of a neutral helium atom =  $1 \times 10^2$  pm. Suppose that we could line up helium atoms side by side in contact with one another. Approximately how many atoms would it take to make the distance from end to end 1 cm?

First, convert 1 cm to picometers.  $1 \text{ cm} \times \frac{0.01 \text{ m}}{1 \text{ cm}} \times \frac{1 \text{ pm}}{1 \times 10^{-12} \text{ m}} = 1 \times 10^{10} \text{ pm}$

$$? \text{ He atoms} = (1 \times 10^{10} \text{ pm}) \times \frac{1 \text{ He atom}}{1 \times 10^2 \text{ pm}} = 1 \times 10^8 \text{ He atoms}$$

- 11 What do we call atoms of the same elements with different mass numbers? Isotopes.

- 13 What is the mass number of an iron atom that has 28 neutrons?

For iron, the atomic number  $Z$  is 26. Therefore the mass number  $A$  is:  $A = 26 + 28 = 54$

- 15 For the following atoms, determine the number of protons and neutrons.

Isotope	${}^3_2\text{He}$	${}^4_2\text{He}$	${}^{24}_{12}\text{Mg}$	${}^{25}_{12}\text{Mg}$	${}^{48}_{22}\text{Ti}$	${}^{79}_{35}\text{Br}$	${}^{195}_{78}\text{Pt}$
No. Protons	2	2	12	12	22	35	78
No. Neutrons	1	2	12	13	26	44	117

- 17 Give the symbol for the following isotopes:

(a)  $Z = 11$   $A = 23$   ${}^{23}_{11}\text{Na}$  (b)  $Z = 28$   $A = 64$   ${}^{64}_{28}\text{Ni}$

- 20 State two differences between a metal and a nonmetal.

Metals conduct heat and electricity, are shiny & malleable; nonmetals are poor conductors, generally dull and brittle.

- 21 Write the names and symbols for four elements in each of the following categories:

(a) nonmetal: phosphorous P, bromine Br, neon, Ne, carbon C

(b) metal: iron Fe, cesium Cs, iridium, Ir, praseodymium Pr

(c) metalloid: silicon Si, germanium Ge, boron B, antimony Sb

- 22 Define, with two examples, the following terms:

(a) Alkali metals: (1<sup>st</sup> column) Sodium, rubidium;

(b) alkaline earth metals: (2<sup>nd</sup> column) barium, beryllium;

(c) halogens: (7<sup>th</sup> column) fluorine, iodine;

(d) noble gases: (8<sup>th</sup> column) xenon, krypton

- 27 What is the difference between an atom and a molecule?

An atom is the smallest representative unit of an element.

A molecule is the smallest representative unit of a (non-ionic) compound, composed of two or more atoms chemically bonded.

- 28 What are allotropes? Allotropes are different physical forms of the same element.

Examples? graphite and diamond are two allotropes of carbon.

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**Michael Brown**



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### **Table of Contents Chemistry Summer Assignment Answer Key**

1. Understanding the eBook Chemistry Summer Assignment Answer Key
  - The Rise of Digital Reading Chemistry Summer Assignment Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Summer Assignment Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Summer Assignment Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Summer Assignment Answer Key
  - Personalized Recommendations
  - Chemistry Summer Assignment Answer Key User Reviews and Ratings
  - Chemistry Summer Assignment Answer Key and Bestseller Lists
5. Accessing Chemistry Summer Assignment Answer Key Free and Paid eBooks
  - Chemistry Summer Assignment Answer Key Public Domain eBooks
  - Chemistry Summer Assignment Answer Key eBook Subscription Services

- Chemistry Summer Assignment Answer Key Budget-Friendly Options
- 6. Navigating Chemistry Summer Assignment Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Summer Assignment Answer Key Compatibility with Devices
  - Chemistry Summer Assignment Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Summer Assignment Answer Key
  - Highlighting and Note-Taking Chemistry Summer Assignment Answer Key
  - Interactive Elements Chemistry Summer Assignment Answer Key
- 8. Staying Engaged with Chemistry Summer Assignment Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Summer Assignment Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Summer Assignment Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Summer Assignment Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Summer Assignment Answer Key
  - Setting Reading Goals Chemistry Summer Assignment Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Summer Assignment Answer Key
  - Fact-Checking eBook Content of Chemistry Summer Assignment Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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