

Chapter 3

1. An 18-g sample of element A combines completely with a 4-g sample of element B to form the compound AB. What is the mass of the compound formed?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_A + \text{Mass}_B = \text{Mass}_{AB}$$

$$\text{Mass}_{AB} = 18 \text{ g} + 4 \text{ g} = 22 \text{ g}$$

2. A substance breaks down into three component elements when it is heated. The mass of each component element is listed in the table below. What was the mass of the substance before it was heated?

Component	Mass (g)
A	39.10
B	54.94
C	64.00

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}} = 39.10 + 54.94 + 64.00 = 158.04 \text{ g}$$

3. Silver iodide powder has been used as an anti-septic and as an agent to seed clouds for rain. Silver iodide is 45.9% silver by mass. If you separate a 50-g sample of silver iodide into its elements, silver and iodine, how much silver would you have?

From the conservation of mass, the mass of silver recovered is equal to the mass of silver in the initial silver iodide sample. The amount of silver recovered would be $50.0 \text{ g} \times 45.9\% = 50.0 \times 0.459 = 22.95 \text{ g} = 23 \text{ g}$.

4. If 5 g of element A combines with 16 g of element B to form compound AB, how many grams of B are needed to form compound AB₂? How many grams of B are needed to form AB₃?

Compound AB₂ contains twice as much element B as does compound AB. Therefore, $2 \times 6 \text{ g} = 12 \text{ g}$.

There is 32 g of B in AB₃. Compound AB₃ contains three times as much element B as does compound AB. Therefore, $3 \times 16 \text{ g} = 48 \text{ g}$. There is 48 g of B in AB₃.

5. During a chemical reaction, 2.445 g of carbon reacts with 3.257 g of oxygen to form carbon monoxide gas. How many grams of carbon monoxide are formed in this reaction?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{carbon}} + \text{Mass}_{\text{oxygen}} = \text{Mass}_{\text{carbon monoxide}}$$

$$2.445 \text{ g} + 3.257 \text{ g} = 5.702 \text{ g}$$

6. Ibuprofen has the chemical formula C₁₃H₁₈O₂. It is 75.69% carbon, 8.80% hydrogen, and 15.51% oxygen. How many mg of carbon does a 200-mg tablet of ibuprofen contain?

$$\text{Mass percentage of an element (\%)} =$$

$$\frac{\text{Mass of element}}{\text{Mass of compound}} \times 100\%$$

$$\text{Mass percentage}_{\text{carbon}} = \frac{\text{Mass}_{\text{carbon}}}{\text{Mass}_{\text{compound}}} \times 100\%$$

$$75.69\% \text{ carbon} = \frac{\text{Mass}_{\text{carbon}}}{200 \text{ mg}} \times 100\%$$

$$75.69\% \text{ carbon} \times \frac{200 \text{ mg}}{100\%} = \text{Mass}_{\text{carbon}}$$

$$\text{Mass}_{\text{carbon in the tablet}} = 151.38 \text{ mg}$$

7. During a chemical reaction, 4.032 g of hydrogen combined with oxygen to form 36.032 g of water. How many grams of oxygen reacted?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{hydrogen}} + \text{Mass}_{\text{oxygen}} = \text{Mass}_{\text{water}}$$

$$4.032 \text{ g} + \text{Mass}_{\text{oxygen}} = 36.032 \text{ g}$$

$$\text{Mass}_{\text{oxygen}} = 36.032 \text{ g} - 4.032 \text{ g} = 32 \text{ g}$$

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Paul A. Bartlett, Judith G. Koch



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1. Understanding the eBook Chemistry Supplemental Problems Answers Solutions
 - The Rise of Digital Reading Chemistry Supplemental Problems Answers Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Supplemental Problems Answers Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Supplemental Problems Answers Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Supplemental Problems Answers Solutions
 - Personalized Recommendations
 - Chemistry Supplemental Problems Answers Solutions User Reviews and Ratings
 - Chemistry Supplemental Problems Answers Solutions and Bestseller Lists
5. Accessing Chemistry Supplemental Problems Answers Solutions Free and Paid eBooks
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 - Chemistry Supplemental Problems Answers Solutions Budget-Friendly Options

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 - Chemistry Supplemental Problems Answers Solutions Enhanced eBook Features
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 - Interactive Elements Chemistry Supplemental Problems Answers Solutions
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Supplemental Problems Answers Solutions
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 - Creating a Diverse Reading Collection Chemistry Supplemental Problems Answers Solutions
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemistry Supplemental Problems Answers Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Supplemental Problems Answers Solutions
 - Fact-Checking eBook Content of Chemistry Supplemental Problems Answers Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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