

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

$$Q \text{ lost (Al)} = Q \text{ gained (Water)}$$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

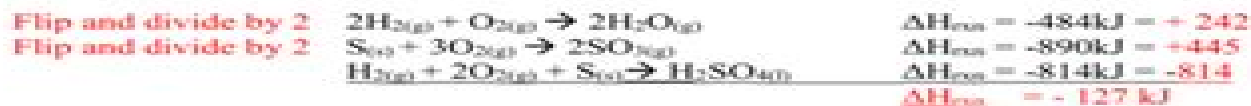
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300 g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3 kJ/mol .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_{4(l)}$ given the following information?



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Table of Contents Chemistry Thermochemistry Test Review Questions

1. Understanding the eBook Chemistry Thermochemistry Test Review Questions
 - The Rise of Digital Reading Chemistry Thermochemistry Test Review Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Thermochemistry Test Review Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Thermochemistry Test Review Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Thermochemistry Test Review Questions
 - Personalized Recommendations
 - Chemistry Thermochemistry Test Review Questions User Reviews and Ratings
 - Chemistry Thermochemistry Test Review Questions and Bestseller Lists
5. Accessing Chemistry Thermochemistry Test Review Questions Free and Paid eBooks
 - Chemistry Thermochemistry Test Review Questions Public Domain eBooks
 - Chemistry Thermochemistry Test Review Questions eBook Subscription Services
 - Chemistry Thermochemistry Test Review Questions Budget-Friendly Options

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 - Chemistry Thermochemistry Test Review Questions Compatibility with Devices
 - Chemistry Thermochemistry Test Review Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Thermochemistry Test Review Questions
 - Highlighting and Note-Taking Chemistry Thermochemistry Test Review Questions
 - Interactive Elements Chemistry Thermochemistry Test Review Questions
8. Staying Engaged with Chemistry Thermochemistry Test Review Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Thermochemistry Test Review Questions
9. Balancing eBooks and Physical Books Chemistry Thermochemistry Test Review Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Thermochemistry Test Review Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Thermochemistry Test Review Questions
 - Setting Reading Goals Chemistry Thermochemistry Test Review Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Thermochemistry Test Review Questions
 - Fact-Checking eBook Content of Chemistry Thermochemistry Test Review Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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