

Chemistry: Unit 10 Solubility Test Review

1. In a solution, the substance being dissolved is the _____.
2. The most commonly used solvent is _____.
3. Ethylene glycol (antifreeze) and water are completely miscible. If you mix 20 mL of ethylene glycol with 10 mL of water in a beaker, what would you expect to see?
4. A solution of sodium chloride and water is what type of mixture?
5. The most precise description of water and oil would be to say that these two substances are _____.
6. Give several examples of an emulsifying agent?
7. What are the properties of an emulsifying agent? (In other words, what do they do? For what are they used?)
8. A silver colored dental filling is a(s) _____ and an iodine solution for cuts is a(n) _____.
9. Give two examples of alloys.
10. What are the four factors that affect the rate of dissolution?

11. Some drugs taken by athletes can show up on drug tests many months after the athlete last took them. These drugs are probably _____ and _____.

Use your solubility curve to answer the next five questions.

12. Between 40°C and 60°C, which compound shows the largest increase in solubility?
13. The solubility of which compound is LEAST affected by temperature?
14. The solubility of KNO_3 increases by approximately _____ when the temperature rises from 10°C to 80°C.
15. For 140 g of KI (in 100 g H_2O) at 15°C, the solution would be described as being _____.
16. Which of the following solutions is unsaturated?
80 g of KCl (in 100 g of water) at 80°C
40 g of NaCl (in 100 g of water) at 50°C
140 g of $NaNO_3$ (in 100 g of water) at 40°C
40 g of H_2SO_4 (in 100 g of water) at 40°C
17. A solution of HCl is 0.380 M. Determine the mass of HCl needed to make 150 mL of solution.
18. A 2.50 M solution of Na_2CO_3 contains 87 g of solute. Find the volume of the solution.

Chemistry Unit 10 Solutions Test Review Answers

**Andrew Burrows, Andrew
Parsons, Gwen Pilling, Gareth Price**

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