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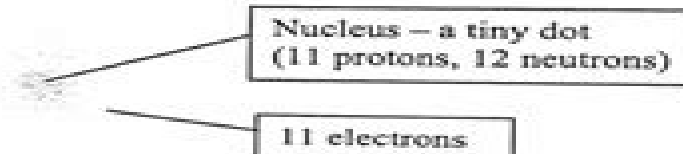
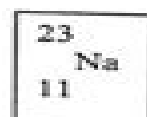
Subatomic Particles

Particle	Symbol	Relative Charge	Absolute Mass	Relative Mass
electron	e^{-}	-1	$9.109 \times 10^{-31} \text{ kg}$	0
proton	p^{+}	+1	$1.673 \times 10^{-27} \text{ kg}$	1
neutron	n^0	0	$1.675 \times 10^{-27} \text{ kg}$	1

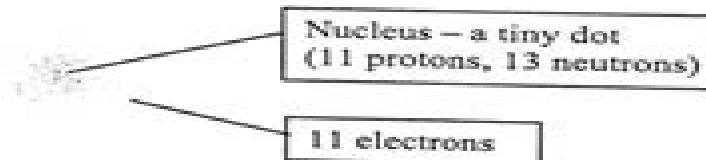
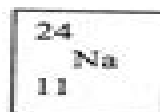
Model: Two Isotopes of Sodium

The diagrams below show representations of sodium isotopes. [Note: the diameter of an atom is about 10,000 times larger than the diameter of the atomic nucleus so the relative sizes of the atom and the nucleus are not accurately depicted in these diagrams.]

Isotope 1



Isotope 2

**Key Questions**

1. What information is provided by the atomic number, Z?

of protons

2. What information is provided by the mass number, A?

protons + neutrons

Chemistry Unit 2 Packet Key

Bowker Editorial Staff



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Table of Contents Chemistry Unit 2 Packet Key

1. Understanding the eBook Chemistry Unit 2 Packet Key
 - The Rise of Digital Reading Chemistry Unit 2 Packet Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 2 Packet Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 2 Packet Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 2 Packet Key
 - Personalized Recommendations
 - Chemistry Unit 2 Packet Key User Reviews and Ratings
 - Chemistry Unit 2 Packet Key and Bestseller Lists
5. Accessing Chemistry Unit 2 Packet Key Free and Paid eBooks
 - Chemistry Unit 2 Packet Key Public Domain eBooks

- Chemistry Unit 2 Packet Key eBook Subscription Services
- Chemistry Unit 2 Packet Key Budget-Friendly Options
- 6. Navigating Chemistry Unit 2 Packet Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 2 Packet Key Compatibility with Devices
 - Chemistry Unit 2 Packet Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 2 Packet Key
 - Highlighting and Note-Taking Chemistry Unit 2 Packet Key
 - Interactive Elements Chemistry Unit 2 Packet Key
- 8. Staying Engaged with Chemistry Unit 2 Packet Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 2 Packet Key
- 9. Balancing eBooks and Physical Books Chemistry Unit 2 Packet Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 2 Packet Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Unit 2 Packet Key
 - Setting Reading Goals Chemistry Unit 2 Packet Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Unit 2 Packet Key
 - Fact-Checking eBook Content of Chemistry Unit 2 Packet Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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