

Name: _____

Date: _____

Period: _____

Chemistry Unit 3 Review

Part 1: History of the Atom

| Scientist/Model name | Drawing/Illustration | Could account for | Could NOT account |
|----------------------|----------------------|---|--|
| Dalton | | <i>All matter is made of atoms. Atoms are small and indestructible. All atoms of a given element are identical in mass and properties</i> | |
| JJ Thomson | | <i>Discovered electrons</i> | <i>Stability of atom</i> |
| Rutherford | | <i>Discovered proton</i> | <i>Stability of atom</i> |
| Bohr | | <i>Energy levels</i> | <i>Continuous and discontinuous spectrum</i> |

Answer the following questions:

Time Line review

Who was the ancient Greek philosopher who first proposed the notion of the atom? *Democritus*

Who's model first introduced the concept of energy levels? *Bohr*

1. What were the major problems of Dalton's atomic theory? *His first two postulates were incorrect.*
2. Whose model was nicknamed the plum pudding model? *Thomson*
3. Which experiment verified that atoms could NOT be solid and the center had a positive charge? *Rutherford*
4. What did Rutherford's gold foil experiment show about the atom? *The nucleus was positively charged.*
5. Rutherford's model of the atom was most similar to Bohr model.
6. What atomic models did not have empty space? *Thomson*

Chemistry Unit 3 Review Answers

The Princeton Review



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