Name

Date ______ Pd ____

Chemistry - Unit 7 Reaction Equations Worksheet 1

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

- 1. $_1_C + _1_{H_2O} \rightarrow _1_{CO} + _1_{H_2}$
- 2. $_2MgO \rightarrow _2Mg + _1_O_2$
- _4_Al + _3_O₂ → _2_Al₂O₃

	#3	Before	After
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- 4. _1_Zn + _1_H₂SO₄ → _1_ZnSO₄ + _1_H₂
- _1_Cl₂ + _2_Kl → _2_KCl + _1_l₂
- _2_CuCl → _2_Cu + _1_Cl₂
- _2_Na + _1_Cl₂ → _2_NaCl
- 8. _2_Al + _6_HCl → _2_AlCl₃ + 3 H₂

- 10. $_4_P + _5_O_2 \rightarrow _2_P_2O_5$
- 11. $_1$ _Mg + $_2$ _HCl \rightarrow $_1$ _MgCl₂ + $_1$ _H₂
- 12. 3 H₂ + 1 N₂ → 2 NH₃
- 13. $_1$ _BaCl₂ + $_1$ _H₂SO₄ \rightarrow $_1$ _BaSO₄ + $_2$ _HCl
- 14. $_1$ _CH₄ + $_2$ _O₂ \rightarrow $_1$ _CO₂ + 2 H₂O

#14 Before After

Before

- 15. a) ___ZnCl₂ + _ NH₄)₂S → ___ZnS + ___ NH₄Cl
 - b) Find the molar mass of these reactants.
 - c) How many moles of ZnCl₂ would be in 25 g? How much mass would 0.55 moles of (NH₄)₂S have?

After

Chemistry Unit 7 Reaction Equations

Xiaolong Qi

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