

QUIZZZ LIVE

AP Chemistry Unit 8 Review
 8/17 questions · 24
 2:00:00:00 · 30:00:00
 @ 10:00:00 · 10:00:00

1.



2.

B.1 Introduction to Acids & Bases

Autoionization of Water:



This is an endothermic process—increasing the heat will shift the reaction to produce more H_3O^+ and OH^- . K_w increases (more products), both pH and pOH will decrease but your H_3O^+ will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$\text{pH} = 3.45$
 $\rightarrow 4 \text{ sig. figs.}$

In case of acidic solution $\rightarrow [\text{H}^+] > [\text{OH}^-]$
 $\text{pH} < 7 \text{ at } 25^\circ\text{C}$
 In case of neutral solution $\rightarrow [\text{H}^+] = [\text{OH}^-]$
 $\text{pH} = 7 \text{ at } 25^\circ\text{C}$
 In case of basic solution $\rightarrow [\text{H}^+] < [\text{OH}^-]$
 $\text{pH} > 7 \text{ at } 25^\circ\text{C}$

ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also know:
 $[\text{H}^+] = 10^{-\text{pH}}$ and $[\text{OH}^-] = 10^{-\text{pOH}}$

3. The value of K_w for water at 0°C is 1×10^{-14} . What is the pOH of water at 0°C ?

- 7.5
 7.0

- 8.0
 6.5

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**Holt Rinehart & Winston, Holt,
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