

Unit 3 Worksheet 3: Mole relationships

Four marks of 1 for each question (total 16)

a. Write the balanced chemical equation

b. Calculate volume of gas at STP when 100g of sodium reacts with sufficient water

c. Use coefficients of the balanced equation to determine mole ratio

d. Which will be the limiting reagent?

1. Hydrogen sulfide gas, which smells like rotten eggs, is formed with the reaction of iron disulfide and water. How many moles of hydrogen gas would be needed to completely react 8 moles of hydrogensulfide with?



2. Ammonia, NH_3 , is formed by reacting hydrogen and nitrogen gases at high temperatures and pressures. How many moles of nitrogen gas is made from 8.0 moles of hydrogen gas?



3. Propane, C_3H_8 , burns in air to form carbon dioxide and water. What moles of carbon dioxide are formed, from 4 moles of propane being burnt?



Chemistry Unit 8 Test B Percent Yield

Max Parsonage



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