

25. Given the reaction:

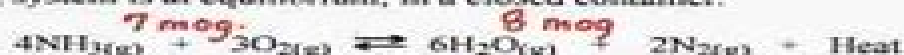


How will the value of the equilibrium constant K_{eq} at 550°C compare with it's value at

450°C ? K_{eq} at 550°C will have a lower value.

Explain your answer. Since the rx. is exothermic, adding heat will cause the reaction to shift LEFT. This will decrease the value of K_{eq} .

26. The following system is at equilibrium, in a closed container:



- a) How is the amount of
- N_2
- in the container affected if the volume of the container is

doubled? Increased (decreasing P will cause a shift RIGHT producing more N_2)

- b) How is the rate of the forward reaction affected if more water vapor is introduced into the container?
- Increase (First inc. $[\text{H}_2\text{O}]$ will speed up rev. rx, forming more reactants. Then forward rx. will speed up.)

- c) How is the amount of
- O_2
- in the container affected if a catalyst is added?

No change (A catalyst speeds up forward & rev. rx. equally)

27. At a certain temperature,
- K_{eq}
- for the reaction:



If the equilibrium concentration of C_2H_2 is 0.40 moles/L, what is the equilibrium concentration of C_6H_6 ?

$$K_{eq} = \frac{[\text{C}_6\text{H}_6]}{[\text{C}_2\text{H}_2]^3} \rightarrow 5.0 = \frac{[\text{C}_6\text{H}_6]}{(0.40)^3} \rightarrow [\text{C}_6\text{H}_6] = (5.0)(0.40)^3 = 0.32 \text{ M}$$

Answer 0.32 M

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