

Chemistry – Unit 5

Worked Example 5 – Gases Again

11. A sealed system (fixed amount) containing gas was investigated. The pressure of the gas at 10°C was 100 kPa. The temperature is 50°C. The total volume of the gas molecules (compared to other gas molecules) is 10%. What is the pressure at the new state?

12. A 100 mL volume of carbon monoxide gas is present at 10°C and 100 kPa. An inert substance reduced the pressure of this gas to 70 kPa at 50°C. Assume temperature and volume constant.

13. A small bubble containing methane is taken where the temperature is 10°C and the pressure is 100 kPa. In the bubble's surface, where the temperature is 50°C and pressure is 10 kPa, calculate the ratio of volume of the bubble at the second volume to the first.

14. Three gases are mixed in a 1.00 L container. The partial pressures of CO_2 is 400 mm Hg, N_2 is 500 mm Hg, and there is 100 mm Hg. What is the pressure of the mixture of gases?

15. What is the percentage by volume of the gases that above mixture?

Chemistry Unit 9 Ws 2 Answers

Yijin Wang

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