

# C2 - THE PERIODIC TABLE 2.6 The transition metals

## GCSE Chemistry Aiming for Grade 4

Where are the transition metals on the periodic table?  
In the centre block

Give the names of five transition metals

**Gold, Silver, Aluminium, Iron, Copper**

List the physical properties of the transition metals

1. Good conductors of heat
2. Good conductors of electricity
3. High densities
4. High melting points
5. Hard and strong

Why is mercury not typical as a transition metal?

It is a liquid at room temperature  
whereas all the others are solid with  
high melting points.



Assessment  
Grade 4

## GCSE Chemistry Aiming for Grade 6

Compared to Group 1 metals how do the following  
properties compare?

**Density**

**Much higher than group 1, none of them will float on  
water like lithium, sodium and potassium.**

**Melting point**

**Much higher melting points except for mercury**

**Boiling point**

**Boiling points are extremely high compared to the  
alkali metals**

**Hardness**

**Alkali metals can be cut with a knife and are soft.  
Transition metals are very hard.**

**Strength**

**Transition metals are very strong and can be used for  
buildings and machines due to their strength**

Give the names of the following transition metal  
compounds.

**$\text{Fe}_2\text{O}_3$**

**Iron Oxide**

**$\text{Cu}_2\text{O}$**

**Copper Oxide**

**$\text{NiCO}_3$**

**Nickel Carbonate**

**$\text{AgNO}_3$**

**Silver Nitrate**

Assessment  
Grade 6

## GCSE Chemistry Aiming for Grade 8

Why is copper used for wires?  
It is a very good conductor -  
other metals. It is also very  
drawn into wires.

Steel is made from Iron and  
make girders on bridges?  
It is very strong and hard to  
weight and high forces.

Titanium is used for aircraft  
Titanium has a very high melting  
point ideal for very hot engine parts

State the colours of the compounds  
of transition metals.

Copper sulphate **blue** due to

Nickel Carbonate **pale green**

Chromium oxide **dark green**

Manganese Chloride **pale pink**

Compared with Group 1 elements  
metals have different properties  
Transition metals can form more than one ion  
their reactivity is different.  
eg Iron can form  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$   
Different reactivity.

Assessment  
Grade 8

# Chemistry Unit C2 Mark Scheme

**Ying Liu**



## **Chemistry Unit C2 Mark Scheme:**

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