

# Transition Metals

## Introduction: Changes of Oxidation States of Transition Elements

By the end of this section, you should be able to predict the charge that various d-block transition metals will have in the periodic table. However, d-block transition elements do not easily distinguish between oxidation states in solution. The d-block transition metal charges:

Group 4g:  $+2$  to  $+3$       Group 5g:  $+2$  to  $+3$       Group 6g:  $+2$  to  $+3$

## Critical Thinking Questions

1. Identify the metal for the following compounds:

$\text{CrCl}_3$

$\text{FeSO}_4$

$\text{Co(NO}_2)_3$

## Introduction: Oxidation States of Transition Elements

Many transition metals can form multiple oxidation states when they form ions. For example, iron can form  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$ . As another example, cobalt can form  $\text{Co}^{2+}$  and  $\text{Co}^{3+}$ .

## Critical Thinking Questions

1. Write the oxidation state of the  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$  ions in the following compounds, showing the steps for computing the correct value. Provide the oxidation state of the metal in each of the following compounds:

$\text{FeCl}_2$

$\text{FeCl}_3$

$\text{Fe}_2\text{O}_3$

$\text{Fe}_2\text{S}_3$

2. Draw and label the steps in writing the compound's correct formula. (They are provided below.)

$\text{FeCl}_2$

$\text{FeCl}_3$

$\text{Fe}_2\text{O}_3$

$\text{Fe}_2\text{S}_3$

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**M. J. Mays**



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## **Table of Contents Chemquest 26 Transition Metals**

1. Understanding the eBook Chemquest 26 Transition Metals
  - The Rise of Digital Reading Chemquest 26 Transition Metals
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemquest 26 Transition Metals
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemquest 26 Transition Metals
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 26 Transition Metals
  - Personalized Recommendations
  - Chemquest 26 Transition Metals User Reviews and Ratings
  - Chemquest 26 Transition Metals and Bestseller Lists
5. Accessing Chemquest 26 Transition Metals Free and Paid eBooks
  - Chemquest 26 Transition Metals Public Domain eBooks
  - Chemquest 26 Transition Metals eBook Subscription Services
  - Chemquest 26 Transition Metals Budget-Friendly Options

6. Navigating Chemquest 26 Transition Metals eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemquest 26 Transition Metals Compatibility with Devices
  - Chemquest 26 Transition Metals Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemquest 26 Transition Metals
  - Highlighting and Note-Taking Chemquest 26 Transition Metals
  - Interactive Elements Chemquest 26 Transition Metals
8. Staying Engaged with Chemquest 26 Transition Metals
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemquest 26 Transition Metals
9. Balancing eBooks and Physical Books Chemquest 26 Transition Metals
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemquest 26 Transition Metals
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemquest 26 Transition Metals
  - Setting Reading Goals Chemquest 26 Transition Metals
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemquest 26 Transition Metals
  - Fact-Checking eBook Content of Chemquest 26 Transition Metals
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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