

balancing Equations

Name: _____

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Introduction: Balancing

- Write the equation for the reaction of sodium and chlorine (chlorine gas) to form sodium chloride.



- Write the equation for the reaction of sodium oxide and carbon dioxide to produce sodium carbonate sodium.



Information: Subscripts and Coefficients

It is important to understand that when you have subscripts in a compound, the number of that subscript tells you the subscript value of that atom. There are no subscript changes for balancing and carbon.

A coefficient tells you how many molecules of a compound there are, that is a different way. For example for the equation: " $2H_2O$ " the 2 is the coefficient. The 2 tells us that there are 2 molecules of water present. In the equation " $1H_2O$ " there are a total of 2 hydrogen atoms and 1 oxygen atom.

Balancing Chemical Equations

- How many oxygen atoms are present in $4H_2O$? $4 \times 1 = 4$
- How many oxygen atoms are in each of the following:



Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride (halogenation):

$$Na + Cl_2 \rightarrow NaCl$$

Remember that when chlorine is the reactant it is always written as Cl_2 . On the product side of the equation, there must always be a total of two atoms of chlorine. Notice the product side there is only one atom of chlorine. Notice we are missing chlorine. In order to fix this equation for public consumption we need to balance the equation. This can be done. But, by adding a "2" to the product side:



Chemquest Balancing Equations Answers

Chris McMullen



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