# Isotopes

The number of protons in a nucleus determines the identity of the element. For example, any atom having 6 protons will be a "carbon" atom. If we were to add an extra proton to the nucleus, we would have an entirely different element. For example,

On the other, hand, if we add an extra NEUTRON to a nucleus we simply end up with the same element, just a little heavier, since the charge on the nucleus would be unchanged.

ISOTOPES of a given element have the same ATOMIC NUMBER but a different ATOMIC MASS.

In other words, isotopes have the same number of protons but a different number of neutrons. An isotope is identified by its mass number, the sum of the protons and neutrons. The most common isotope of Carbon has a mass number of 12 and can be written as Carbon-12, two other isotopes are Carbon-13 and Carbon-14. Despite their different mass numbers, all three carbon isotopes react the same way chemically.

#### PART I. Answer the questions based on the above reading.

 What is an isotope? Isotopes are versions of the same element. They have the same number of protons and electrons as the element but different mass numbers and number of neutrons.

What does the number next to isotopes signify? The number indicates the isotope's mass number.

How can you tell isotopes of the same element apart? They will have a different mass number and different number of neutrons.

# PART II. For each of the following isotopes, write the number of protons, neutrons, and electrons. Assume all atoms are neutral.

	Chromium-58	Chromium-63
# of protons	24	24
# of neutrons	34	39
# of electrons	24	24

	Carbon- 12	Carbon- 13	Carbon- 14
# of protons	6	6	6
# of neutrons	6	7	8
# of electrons	6	6	6

# **Chemthink Answer Key Isotopes**

**B** Lingard

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