

## 4 • Chemical Equations and Stoichiometry

### SIX STOICHIOMETRY PROBLEMS

#### General Stoichiometry

13. Several brands of antacid tablets use aluminum hydroxide to neutralize excess acid.



[Molar masses:            78.01            36.46            133.4            18.02]

What quantity of HCl, in grams, can a tablet with 0.750 g of  $\text{Al}(\text{OH})_3$  consume? What quantity of water is produced?

**1.05 g HCl**

**0.520 g  $\text{H}_2\text{O}$**

17. The equation for one of the reactions in the process of reducing iron ore to the metal is



[Molar masses:            159.7            28.01            55.85            44.01]

- (a) What is the maximum mass of iron, in grams, that can be obtained from 454 g (1.00 lb) of iron(III) oxide?

**31.8 g Fe**

- (b) What mass of CO is required to reduce the iron(III) oxide to iron metal?

**239 g CO**

#### Limiting Reactants

23. The reaction of methane and water is one way to prepare hydrogen:



[Molar masses:            16.04            18.02            28.01            2.02]

If you begin with 995 g of  $\text{CH}_4$  and 2510 g of water, what is the maximum possible yield of  $\text{H}_2$ ?

**$\text{CH}_4$  is LR.... 376 g  $\text{H}_2$  is formed**

# Chemthink Answers Stoichiometry

**David E. Newton**



## **Chemthink Answers Stoichiometry:**

**Stoichiometry** Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay's Phase rule which was the first volume of the series issued. It belongs properly however to this volume and is therefore included here. Note p vii Introduction issued also separately in 1904 Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry. As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all. The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities. For the serious student this means more often than not the frustrations of questions unanswered. The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer. Distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media. Quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions. Intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction. You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers. Covers calculations relating to solids, solutions, gases and electrolysis plus as limiting and excess reactants, chemical yields, atom economy and much more. Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018. A Stoichiometry Unit David Callaghan, 2004 **Stoichiometry and structure** Michell J. Sienko, 1964

**Chemistry Problems** David E. Newton, 2001 This edition includes acid base chemistry and thermochemistry. Chemistry Problems is the authoritative resource for practice problems covering all the essentials. Includes Atomic structure, Stoichiometry, Solutions chemistry and Electrochemistry. Literally thousands of problems in this compendium build proficiency, analytical skills and math skills. The text includes a complete answer key and reference to applicable web sites.

**Stoichiometry** B. I. Bhatt, 1996 Bite-Sized Chemistry Calculations John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension. Written by an experienced chemistry educator, each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide. This part of the series explores the different types of problems and calculations.

encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance

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Volume Stoichiometry

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