

## Chemthink: CHEMICAL REACTIONS

- a) Go to [www.chemthink.com](http://www.chemthink.com). Log in using your user name and password. (If you forgot, you need to WAIT QUIETLY until your teacher is done introducing the lesson.)
- b) Scroll to **CHEMICAL REACTIONS** then the tutorial under the Chemical Reactions menu.
- c) Complete the **tutorial**. As you complete the tutorial, write down the answers to the questions below.
- d) Once you have successfully completed the tutorial and this worksheet, complete the **Question Set**.

### CHEMICAL REACTIONS TUTORIAL QUESTIONS:

- Starting materials in a chemical reaction are called \_\_\_\_\_.
- The ending materials in a chemical reaction are called \_\_\_\_\_.
- The arrow indicates a \_\_\_\_\_ has taken place.
- All reactions have one thing in common: there is a \_\_\_\_\_ of chemical bonds.
- Chemical reactions always involve \_\_\_\_\_, old bonds, \_\_\_\_\_, new bonds, or both.
- In all reactions we still have all of the \_\_\_\_\_ at the end that we had at the start.
- In every reaction there can never be any \_\_\_\_\_ atoms or \_\_\_\_\_.
- Chemical reactions only \_\_\_\_\_ in the atoms that are already there.
- Let's represent a reaction on paper. For example, hydrogen gas ( $H_2$ ) reacts with oxygen gas ( $O_2$ ) to form water ( $H_2O$ ):



If we use only the atoms shown, we'd have \_\_\_\_\_ atoms of H and \_\_\_\_\_ atoms of O as reactants. This would make \_\_\_\_\_ molecule of  $H_2O$ , but we'd have \_\_\_\_\_ atom of O leftover. However, this reaction only makes  $H_2O$ .

Remember: reactions are not limited to 1 molecule each of reactants. We can use as many as we need to balance the chemical equation.

#### A balanced chemical reaction shows:

- a) What atoms are present before (in the reactants) and after (in the products)
  - b) **How many** of each reactant and product is present before and after.
10. So to make  $H_2O$  from oxygen gas and hydrogen gas, the balanced equation would be:



Which is the same as:

# of atoms in Reactants	Element	# of atoms in Products
	H	
	O	

- This idea is called the \_\_\_\_\_.
- There must be the same \_\_\_\_\_ and the same number of \_\_\_\_\_ before the reaction (in the reactants) and after the reaction (in the products).
- What is the balanced equation for this reaction? \_\_\_\_\_  $Cu$  + \_\_\_\_\_  $O_2 \rightarrow$  \_\_\_\_\_  $CuO$

# Chemthink Chemical Reactions Questions Answers

**John Joseph Griffin**



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## **Table of Contents Chemthink Chemical Reactions Questions Answers**

1. Understanding the eBook Chemthink Chemical Reactions Questions Answers
  - The Rise of Digital Reading Chemthink Chemical Reactions Questions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemthink Chemical Reactions Questions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemthink Chemical Reactions Questions Answers
  - Personalized Recommendations
  - Chemthink Chemical Reactions Questions Answers User Reviews and Ratings
  - Chemthink Chemical Reactions Questions Answers and Bestseller Lists

5. Accessing Chemthink Chemical Reactions Questions Answers Free and Paid eBooks
  - Chemthink Chemical Reactions Questions Answers Public Domain eBooks
  - Chemthink Chemical Reactions Questions Answers eBook Subscription Services
  - Chemthink Chemical Reactions Questions Answers Budget-Friendly Options
6. Navigating Chemthink Chemical Reactions Questions Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemthink Chemical Reactions Questions Answers Compatibility with Devices
  - Chemthink Chemical Reactions Questions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemthink Chemical Reactions Questions Answers
  - Highlighting and Note-Taking Chemthink Chemical Reactions Questions Answers
  - Interactive Elements Chemthink Chemical Reactions Questions Answers
8. Staying Engaged with Chemthink Chemical Reactions Questions Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemthink Chemical Reactions Questions Answers
9. Balancing eBooks and Physical Books Chemthink Chemical Reactions Questions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemthink Chemical Reactions Questions Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemthink Chemical Reactions Questions Answers
  - Setting Reading Goals Chemthink Chemical Reactions Questions Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemthink Chemical Reactions Questions Answers
  - Fact-Checking eBook Content of Chemthink Chemical Reactions Questions Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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