

(vi) In the reaction $I_2 + I^- \longrightarrow I_3^-$, which is the Lewis base?

(a) I_2

(b) I^-

(c) I_3^-

(d) None of these

(vii) In what manner will increase of pressure affect the following equilibrium?



(a) No effect

(b) Shift in the forward direction

(c) Shift in the reverse direction

(d) Increase in the field of H_2

(viii) A Metal ion M^{3+} loses 3 electrons, its oxidation number will be

(a) +3

(b) -3

(c) 0

(d) +6

(ix) The oxidation states of iodine in HIO_4 , H_3IO_5 and H_5IO_6 are respectively

(a) +1, +3, +7

(b) +7, +7, +3

(c) +7, +7, +7

(d) +7, +5, +3

(x) $:CCl_2$ is

(a) an electrophile

(b) a free radical

(c) a nucleophile

(d) none of these.

SECTION B

Q 2. Answer the following questions.

(i) Write the general electronic configuration of p and d block elements.

(ii) At $0^\circ C$, ice and water are present in equilibrium. What will happen on increasing the pressure.

(iii) Define oxidation and reduction according to classical concept.

(iv) What are Carbocations and Free radicals?

(v) Give two methods of preparation of Alkanes?

(vi) Why is benzene extraordinarily stable though it contains three double bonds?

(vii) Define Inductive effect & Hyper- conjugation.

(viii) Define

a) Resonance

b) Ozonolysis

(ix) Define Sulphonation. Give Example.

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Table of Contents Chemistry Test Sections 11 1 And 11

1. Understanding the eBook Chemistry Test Sections 11 1 And 11
 - The Rise of Digital Reading Chemistry Test Sections 11 1 And 11
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Test Sections 11 1 And 11
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Test Sections 11 1 And 11
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Test Sections 11 1 And 11
 - Personalized Recommendations
 - Chemistry Test Sections 11 1 And 11 User Reviews and Ratings
 - Chemistry Test Sections 11 1 And 11 and Bestseller Lists
5. Accessing Chemistry Test Sections 11 1 And 11 Free and Paid eBooks
 - Chemistry Test Sections 11 1 And 11 Public Domain eBooks

- Chemistry Test Sections 11 1 And 11 eBook Subscription Services
- Chemistry Test Sections 11 1 And 11 Budget-Friendly Options
- 6. Navigating Chemistry Test Sections 11 1 And 11 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Test Sections 11 1 And 11 Compatibility with Devices
 - Chemistry Test Sections 11 1 And 11 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Test Sections 11 1 And 11
 - Highlighting and Note-Taking Chemistry Test Sections 11 1 And 11
 - Interactive Elements Chemistry Test Sections 11 1 And 11
- 8. Staying Engaged with Chemistry Test Sections 11 1 And 11
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Test Sections 11 1 And 11
- 9. Balancing eBooks and Physical Books Chemistry Test Sections 11 1 And 11
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Test Sections 11 1 And 11
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Test Sections 11 1 And 11
 - Setting Reading Goals Chemistry Test Sections 11 1 And 11
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Test Sections 11 1 And 11
 - Fact-Checking eBook Content of Chemistry Test Sections 11 1 And 11
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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