

## Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol,  $\Delta H_{\text{comb}}$



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone,  $\text{CaCO}_3$ , to lime,  $\text{CaO}$ , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{ CaCO}_3(s) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CaO}(s) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CO}_2(g) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4.  $\text{CH}_4(g) + 3 \text{Cl}_2(g) \rightarrow \text{CHCl}_3(l) + 3 \text{HCl}(g) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide,  $\text{CH}_2\text{N}_2$ , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation:  $\Delta H^\circ_f \text{ NO}(g) = 90.3 \text{ kJ/mol}$  and  $\Delta H^\circ_f \text{ H}_2\text{O}(g) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction:  $2 \text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + 2 \text{Fe}(s) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess  $\text{Fe}_2\text{O}_3$ , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

# Chemistry Thermochemistry Practice Problems Answers

**Darrell Ebbing**



## **Chemistry Thermochemistry Practice Problems Answers:**

**Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields. [Chemistry Workbook For Dummies with Online Practice](#) Chris Hren, Peter J.

Mikulecky, 2017-04-17 Take the confusion out of chemistry with hundreds of practice problems. Chemistry Workbook For Dummies is your ultimate companion for introductory chemistry at the high school or college level. Packed with hundreds of practice problems, this workbook gives you the practice you need to internalize the essential concepts that form the foundations of chemistry. From matter and molecules to moles and measurements, these problems cover the full spectrum of topics you'll see in class, and each section includes key concept review and full explanations for every problem to quickly get you on the right track. This new third edition includes access to an online test bank where you'll find bonus chapter quizzes to help you test your understanding and pinpoint areas in need of review. Whether you're preparing for an exam or seeking a start to finish study aid, this workbook is your ticket to acing basic chemistry. Chemistry problems can look intimidating; it's a whole new language with different rules, new symbols, and complex concepts. The good news is that practice makes perfect, and this book provides plenty of it with easy-to-understand coaching every step of the way. Delve deep into the parts of the periodic table. Get comfortable with units, scientific notation, and chemical equations. Work with states, phases, energy, and charges. Master nomenclature, acids, bases, titrations, redox reactions, and more. Understanding introductory chemistry is critical for your success in all science classes, so following up with the material now makes life much easier down the education road. Chemistry Workbook For Dummies gives you the practice you need to succeed. [Chemistry Workbook For Dummies](#) Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases, work chemistry equations and use formulas with ease. Got a grasp on the chemistry terms and concepts you

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