

Energy and Chemical Change

Section 16.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- | | |
|-------|--|
| _____ | 1. <i>Energy</i> is the ability to do work or produce heat. |
| _____ | 2. The law of conservation of energy states that energy <i>can be</i> created and destroyed. |
| _____ | 3. Chemical potential energy is energy stored in a substance because of its <i>composition</i> . |
| _____ | 4. <i>Heat</i> is a form of energy that flows from a warmer object to a cooler object. |
| _____ | 5. A calorie is the amount of energy required to raise the temperature of <i>one gram</i> of pure water by one degree Celsius. |
| _____ | 6. A <i>calorie</i> is the SI unit of heat and energy. |
| _____ | 7. The <i>specific heat</i> of a substance is the amount of heat required to raise the temperature of <i>one gram</i> of that substance by one degree Celsius. |
| _____ | 8. <i>Kinetic energy</i> is energy of motion. |
| _____ | 9. Chemicals participating in a chemical reaction contain <i>only potential energy</i> . |
| _____ | 10. One nutritional Calorie is equal to <i>100 calories</i> . |
| _____ | 11. One calorie equals <i>4.184 joules</i> . |
| _____ | 12. When a fuel is burned, some of its <i>chemical potential energy</i> is lost as heat. |
| _____ | 13. To convert kilojoules to joules, <i>divide</i> the number of kilojoules by 1000 joules/1 kilojoule. |

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J(g·°C).

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McGraw Hill



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Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook** McGraw Hill, 2001-08

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Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition McGraw Hill, 2001-03-30 Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats There are worksheets and study tools for each section of the text that help teachers track students progress toward understanding concepts Guided Reading Activities help students identify and comprehend the important information in each chapter

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Table of Contents Chemistry Study Guide For Content Mastery Solutions

1. Understanding the eBook Chemistry Study Guide For Content Mastery Solutions
 - The Rise of Digital Reading Chemistry Study Guide For Content Mastery Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide For Content Mastery Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide For Content Mastery Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide For Content Mastery Solutions
 - Personalized Recommendations
 - Chemistry Study Guide For Content Mastery Solutions User Reviews and Ratings
 - Chemistry Study Guide For Content Mastery Solutions and Bestseller Lists
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8. Staying Engaged with Chemistry Study Guide For Content Mastery Solutions
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide For Content Mastery Solutions
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- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide For Content Mastery Solutions
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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- Setting Reading Goals Chemistry Study Guide For Content Mastery Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide For Content Mastery Solutions
- Fact-Checking eBook Content of Chemistry Study Guide For Content Mastery Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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