

## KS3 - Periodic Table

On the periodic table below, shade in the metals and non metals in different colours.  
Label group 1, 7 and 0.

The periodic table displays elements from Hydrogen (H) to Oganesson (Og). Elements are color-coded: noble gases (yellow), halogens (green), chalcogens (blue), transition metals (purple), and main group elements (pink). The lanthanide and actinide series are shown at the bottom.

Group 1 metals react violently with water. Watch the demonstration and then answer the questions.

Which metal was the most reactive?

Which metals was the least reactive?

Which colour did potassium burn?

Which colour did sodium burn?

**Group 7 - The Halogens**

Answer the following questions.

- |   |    |
|---|----|
| 1. Which element is the most reactive?        | F  |
| 2. Which element is the least reactive?       | Cl |
| 3. Which element do we use as a bleach?       | Br |
| 4. Which element do we use as a disinfectant? | I  |
|   | At |

F
Cl
Br
I
At

How do we know if something is a metal?

How do we know if something is a non-metal?

What does a good conductor do?

Fill in the missing words.

- The vertical columns of the Periodic Table are  
g\_\_\_\_\_
- The horizontal rows are  
p\_\_\_\_\_
- There are patterns in element properties  
d\_\_\_\_\_ groups and  
a\_\_\_\_\_ periods.

Group 0 - The Noble Gases

List the elements in Group 0

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Noble gases are very unreactive and are used in helium balloons, neon lights and argon gas between our double glazing.

## Displacement reactions

In a chemical reaction a more reactive Group 7 element will push out a less reactive Group 7 element.

For example:  
chlorine + potassium iodide  $\rightarrow$  potassium chloride + iodine

# Chemistry Test Papers Ks3

**Kath Jordan**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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