

Part A: Multiple Choice. Circle the best answer for each question. Answer all questions in this section. (24 pts)

1. A heated 250 g sample of aluminium metal is placed in water at room temperature. The water absorbs 264 J of heat from the metal and increases in temperature to 34.9°C . What was the initial temperature of the aluminium? ($c_{\text{Al}} = 0.900 \text{ J/g}^{\circ}\text{C}$)

a. 11.7°C
b. 23.2°C
c. 36.1°C
d. 46.6°C

2. Given in the thermochemical equation below, what is the molar enthalpy of formation for N_2O ?

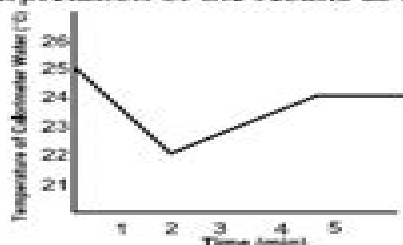


a. -163.3 kJ/mol
b. -81.6 kJ/mol
c. 81.6 kJ/mol
d. 163.3 kJ/mol

3. Which of the following is has an enthalpy of formation of 0 kJ/mol ?

a. $\text{N}_2(\text{g})$
b. $\text{Hg}(\text{s})$
c. $\text{N}_2(\text{g})$
d. $\text{N}_2(\text{l})$

4. A piece of ice was dropped into liquid water to determine the molar heat of fusion of water. What is the best interpretation of the results as shown in the graph below?



a. The ice finishes melting at 2 minutes and heat enters the container
b. The ice finishes melting at 2 minutes and heat leaves the container
c. The ice finishes melting at 5 minutes and heat enters the container
d. The ice finishes melting at 5 minutes and heat leaves the container.

Chemistry Thermochemistry Test Test Answers

Claire Aldridge, Karl Lee



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