

## Chemistry Unit 3 Review

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1. Define the Law of Conservation of Mass

**Mass is not created or destroyed.**

2. Using your knowledge of the types of chemical reaction, balance the following equations and indicate which of the six types of reaction they are:



Type of reaction: **SR**

Redox: **Yes/No**



Type of reaction: **S**

Redox: **Yes/No**



Type of reaction: **D**

Redox: **Yes/No**



Type of reaction: **DR**

Redox: **Yes/No**



Type of reaction: **SR**

Redox: **Yes/No**



Type of reaction: **C**

Redox: **Yes/No**



Type of reaction: **D**

Redox: **Yes/No**

3. Define oxidation-reduction (redox) reaction:

**The charges change from the reactant to product side of the reaction. Oxidation loses electrons- it gets more positive. Reduction gains electrons- it gets more negative.**

4. Given the following word equations, write the complete balanced chemical equation and identify the type of reaction.

- a) Copper (II) chloride reacts with iron to produce iron (III) chloride and copper metal.



- b) Hydrogen gas and bromine liquid react to yield hydrogen bromide.



- c) In a Bunsen burner, carbon tetrahydride (methane gas) reacts with oxygen to produce carbon dioxide and water vapor.



# Chemistry Unit 3 Review V2 2 Answers

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### **Table of Contents Chemistry Unit 3 Review V2 2 Answers**

1. Understanding the eBook Chemistry Unit 3 Review V2 2 Answers
  - The Rise of Digital Reading Chemistry Unit 3 Review V2 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 3 Review V2 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit 3 Review V2 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 3 Review V2 2 Answers
  - Personalized Recommendations
  - Chemistry Unit 3 Review V2 2 Answers User Reviews and Ratings
  - Chemistry Unit 3 Review V2 2 Answers and Bestseller Lists
5. Accessing Chemistry Unit 3 Review V2 2 Answers Free and Paid eBooks

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- ePub, PDF, MOBI, and More
  - Chemistry Unit 3 Review V2 2 Answers Compatibility with Devices
  - Chemistry Unit 3 Review V2 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Unit 3 Review V2 2 Answers
  - Highlighting and Note-Taking Chemistry Unit 3 Review V2 2 Answers
  - Interactive Elements Chemistry Unit 3 Review V2 2 Answers
8. Staying Engaged with Chemistry Unit 3 Review V2 2 Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Unit 3 Review V2 2 Answers
9. Balancing eBooks and Physical Books Chemistry Unit 3 Review V2 2 Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Unit 3 Review V2 2 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit 3 Review V2 2 Answers
- Setting Reading Goals Chemistry Unit 3 Review V2 2 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit 3 Review V2 2 Answers
- Fact-Checking eBook Content of Chemistry Unit 3 Review V2 2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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