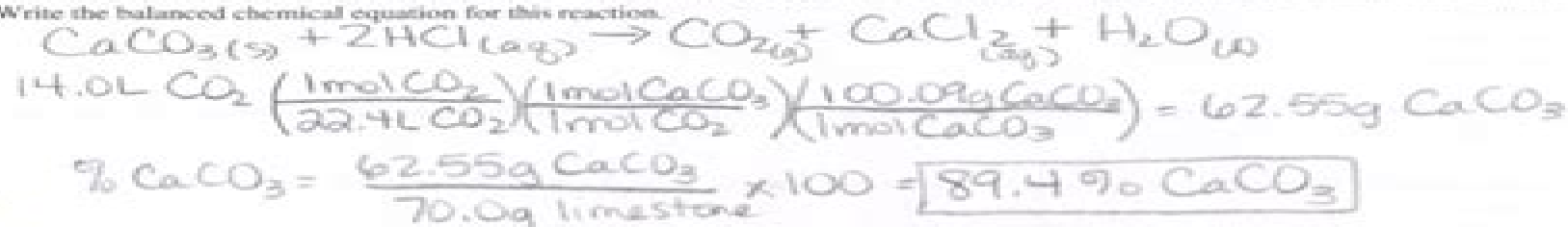
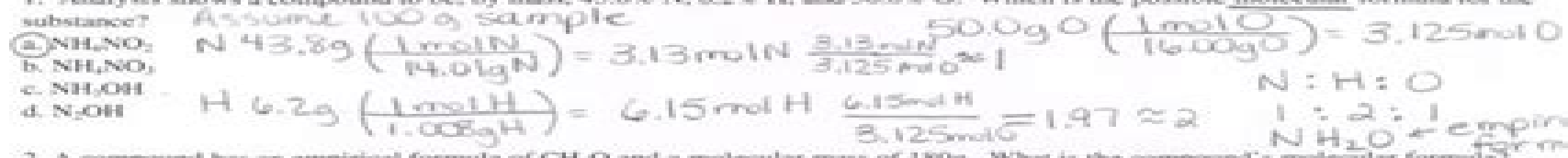


11. A 70.0 g sample of limestone consists of a large percentage of calcium carbonate. The sample reacts with an excess of hydrochloric acid and 14.0 L of carbon dioxide is generated at STP. What is the percentage of calcium carbonate in the limestone? Write the balanced chemical equation for this reaction.

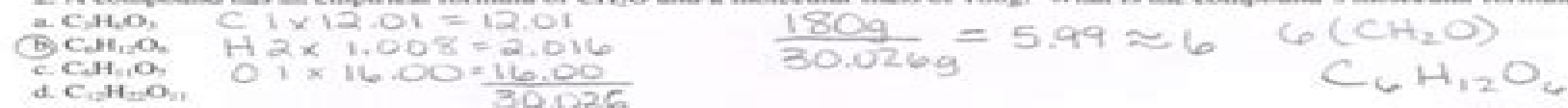


2.2.5 Analyze quantitatively the composition of a substance (empirical formula, molecular formula, percent composition, anhydrides).

1. Analysis shows a compound to be, by mass, 43.8% N, 6.2% H, and 50.0% O. Which is the possible molecular formula for the substance? Assume 100 g sample



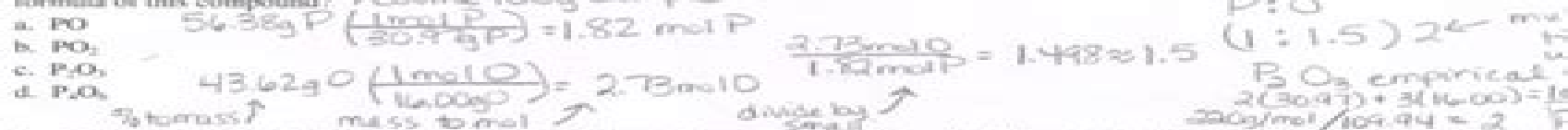
2. A compound has an empirical formula of CH_2O and a molecular mass of 180g. What is the compound's molecular formula?



3. What is the percent by mass of iron in the compound Fe_2O_3 ?



4. A compound consisting of 56.38% phosphorus and 43.62% oxygen has a molecular mass of 220 g/mole. What is the molecular formula of this compound? Assume 100g sample



5. A 10.10 g sample of barium chloride hydrate is heated in a crucible. After all of the water is driven off, 8.50 g of the anhydrous barium chloride remains in the crucible. What is the formula of the hydrate?



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Wolfgang Guggemos

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