Unit 8 - Acids and Bases (11-15%)

Calculate the values of pH and pOH, based on Kw and the concentration of all species present in a neutral solution of water.

8.1 Intro to Arids and Bases: Calculate the values of pH and pOH, based on Kw and the concentration of all species present in a neutral solution of water.

The concentrations of hydronium ion and hydroside ion are often reported as oil and pOH, respectively.

$$pH = -log(H_0O')$$

 $pOH = -log(OH')$

The terms "hydrogen ion" and "hydronium ion" and the symbols H" (ag) and H₂O' (ag) are often used interchangeably for the aqueous ion of hydrogen. Hydronium ion and HyO' (aq) are preferred, but H+ facil is also accepted on the AP Exam.

Water autoionizes with an equilibrium constant Kw.

- In pure water, pit # pOH is called a neutral solution. At 25°C, pixw # 14.0 and thus pit # pOH # 7.0. $\mu E_{rel} = 2.4 + \mu E I + \mu O I I + r 2.5 ^{\circ} C$
- The value of the it temperature dependent, so the priof pure, reutral enter will deviate from 7.0 at

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