AP Chemistry - Unit 9 Practice Test (Thermochemistry and Electrochemistry - Chapters 19 and 20)

 The value of ΔG° at 141.0 °C for the formation of phosphorous trichloride from its constituent elements.

1)

4)

5)

6)

 $P_2(g) + 3Cl_2(g) \rightarrow 2PCl_3(g)$ 

is \_\_\_\_\_ kJ/mol. At 25.0°C for this reaction,  $\Delta H^o$  is -720.5 kJ/mol,  $\Delta G^o$  is -642.9 kJ/mol, and  $\Delta S^o$  is -263.7 J/K.

A1 = 683.3

B)  $3.65 \times 10^4$ 

C) = 829.7

D) -612.3

E)  $1.08 \times 10^5$ 

2) How many grams of copper will be plated out by a current of 2.3 A applied for 25 minutes to a 0.50 2) \_\_\_\_\_\_ M solution of copper(II) sulfate?

A)  $1.8 \times 10^{-2}$ 

B) 2.2

C) 0.019

D) 0.036

E) 1.1

3) The standard cell potential (E\*) of a voltaic cell constructed using the cell reaction below is 0.76 V: 3)

$$Zn(s) + 2H^*(aq) \rightarrow Zn^{2*}(aq) + H_2(g)$$

With  $P_{H_2} = 1.0$  atm and  $[Zn^2+] = 1.0$  M, the cell potential is 0.66 V. The concentration of H+ in the cathode compartment is \_\_\_\_\_ M.

A)  $1.4 \times 10^{-1}$ 

B)  $4.9 \times 10^{1}$ 

C) 4.2 x 10-4 D) 1.0 x 10-12

E)  $2.0 \times 10^{-2}$ 

ΔS is negative for the reaction \_\_\_\_\_\_

A) H<sub>2</sub>O (I) → H<sub>2</sub>O (g)

B) 2SO<sub>2</sub> (g) + O<sub>2</sub> (g) → 2SO<sub>3</sub> (g)

C)  $NH_4Cl(s) \rightarrow NH_3(g) + HCl(g)$ 

D) 2C (s) +  $O_2$  (g)  $\rightarrow$  2CO<sub>2</sub> (g)

E) PbCl<sub>2</sub> (s)  $\rightarrow$  Pb<sup>2+</sup> (aq) + 2Cl<sup>-</sup> (aq)

5) The standard emf for the cell using the overall cell reaction below is +2.20 V:

 $2AI(s) + 3I_2(s) \rightarrow 2AI^{3+}(aa) + 6I^{-}(aa)$ 

The emf generated by the cell when  $[AI^{3+}] = 4.5 \times 10^{-3}$  M and  $[I^{-}] = 0.15$  M is \_\_\_\_\_\_V. A) 2.23 B) 2.32 C) 2.20 D) 2.10 E) 2.39

6) Which substance is the reducing agent in the reaction below?

Pb + PbO<sub>2</sub> + 2H<sub>2</sub>SO<sub>4</sub> → 2PbSO<sub>4</sub> + 2H<sub>2</sub>O

A) H<sub>2</sub>SO<sub>4</sub>

B) PbO<sub>2</sub>

C) Pb

D) H<sub>2</sub>O

E) PbSO<sub>4</sub>

# **Chemistry Unit 9 Thermochemistry Test Answers**

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#### **Table of Contents Chemistry Unit 9 Thermochemistry Test Answers**

- 1. Understanding the eBook Chemistry Unit 9 Thermochemistry Test Answers
  - The Rise of Digital Reading Chemistry Unit 9 Thermochemistry Test Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Unit 9 Thermochemistry Test Answers
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit 9 Thermochemistry Test Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Unit 9 Thermochemistry Test Answers
  - Personalized Recommendations
  - Chemistry Unit 9 Thermochemistry Test Answers User Reviews and Ratings

- Chemistry Unit 9 Thermochemistry Test Answers and Bestseller Lists
- 5. Accessing Chemistry Unit 9 Thermochemistry Test Answers Free and Paid eBooks
  - Chemistry Unit 9 Thermochemistry Test Answers Public Domain eBooks
  - Chemistry Unit 9 Thermochemistry Test Answers eBook Subscription Services
  - Chemistry Unit 9 Thermochemistry Test Answers Budget-Friendly Options
- 6. Navigating Chemistry Unit 9 Thermochemistry Test Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chemistry Unit 9 Thermochemistry Test Answers Compatibility with Devices
  - Chemistry Unit 9 Thermochemistry Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Unit 9 Thermochemistry Test Answers
  - Highlighting and Note-Taking Chemistry Unit 9 Thermochemistry Test Answers
  - Interactive Elements Chemistry Unit 9 Thermochemistry Test Answers
- 8. Staying Engaged with Chemistry Unit 9 Thermochemistry Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Unit 9 Thermochemistry Test Answers
- 9. Balancing eBooks and Physical Books Chemistry Unit 9 Thermochemistry Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Unit 9 Thermochemistry Test Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Unit 9 Thermochemistry Test Answers
  - Setting Reading Goals Chemistry Unit 9 Thermochemistry Test Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Unit 9 Thermochemistry Test Answers
  - Fact-Checking eBook Content of Chemistry Unit 9 Thermochemistry Test Answers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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