

Stoichiometry Worksheet 2 Answer Key

1.

- a. 2 / 13
- b. 13 / 8
- c. 13 / 10
- d. 2 / 8 (or 1 / 4)
- e. 2 / 10 (or 1 / 5)

2. The $\text{KClO}_3 / \text{O}_2$ molar ratio is 2/3.

$$2 \text{ mol KClO}_3 / 3 \text{ mol. O}_2 = 12.00 \text{ mol KClO}_3 / x = 18.00 \text{ mol.}$$

$$x = 18.00 \text{ mol of O}_2$$

3. Given the following equation: $2 \text{ K} + \text{Cl}_2 \rightarrow 2 \text{ KCl}$

How many grams of KCl is produced from 2.50 g of K and excess Cl_2 .

2.50 g of K		=	0.0639 mol of K
39.098 g/mol			
K	2		0.0639 mol of K
KCl	2	=	x
			x = 0.0639 mol of KCl
0.0639 mol x 74.551 g/mol of KCl = 4.77 g of KCl			

From 1.00 g of Cl_2 and excess K?

1.00 g of Cl_2		=	0.0141 mol of Cl_2
70.906 g/mol			
Cl_2	1		0.0141 mol of Cl_2
KCl	2	=	x
			x = 0.0282 mol of KCl
0.0282 mol x 74.551 g/mol = 2.10 g of KCl			

Chemistry Unit 9 Ws 2 Answers Stoichiometry

John Michels (Journalist)



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