

Name Key Class \_\_\_\_\_ Date \_\_\_\_\_

## Nuclear Decay

### Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, positron, and gamma radiation.

Name	Alpha	Beta	Positron	Gamma
Greek letter	$\alpha$	$\beta^-$	$\beta^+$	$\gamma$
Symbol	${}^4_2\text{He}$	${}^0_{-1}e$	${}^0_{+1}e$	N/A
Composition	$2p^+$ , $2n^0$	electron	anti-electron	EM wave
Charge	$2+$	$1-$	$1+$	$0$
Stopped by	paper, skin, clothing	wood, glass	wood, glass	thick concrete or lead

### Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A. alpha emission



B. beta emission



C. electron capture



D. positron emission

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha ( $\alpha$ ) emission.



7. Decay of iodine-126 by positron ( $\beta^+$ ) emission.



8. How are the atomic number and mass number of a nuclide affected by positron emission?

Atomic number decreases by one. Mass number stays the same.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

Both atomic number and mass number stay the same since gamma radiation has no mass or charge.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

The reactants have a slightly greater mass. In a nuclear reaction, a small amount of mass is converted to energy according to the equation  $E = mc^2$ . The difference in mass is referred to as the mass defect.

# Chemistry Worksheet Questions 1 On Radioactivity Answers

**Karin Nielsen-Saines**



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