

Intro to Acids and Bases

Name: _____
 Class: _____
 Date: _____

Define each term: Definitions of acids and bases:

Arrhenius Definitions:

- acid substances that release dissolved hydrogen ions (H^+) (acids H^+ gases formed in gases as the hydrogens form H_2 and acid H^+ form H_2O) and conjugated anions (A^-)
- base substances which release hydroxide ions (OH^-)

Brønsted-Lowry Definitions:

- acid substances that donate a proton (H^+) from another
- base substances which accept a proton (H^+) from another

Table 1: Representative reactions for H^+ (acid) and OH^- (base) equilibrium reactions:

Reaction	K_c
1. $HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$	1.0×10^6
2. $HNO_3(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + NO_3^-(aq)$	2.0×10^5
3. $H_2SO_4(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + HSO_4^-(aq)$	1.0×10^3
4. $H_2SO_4(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + HSO_4^-(aq)$	1.0×10^3
5. $H_2O(l) + H_2O(l) \rightleftharpoons H_3O^+(aq) + OH^-(aq)$	1.0×10^{-14}

Critical Thinking Questions

- Are all the reactions in Table 1 that are equilibrium acids?
- Are all the reactions in Table 1 that are equilibrium bases?

Class work:

- Are all the reactions in Table 1 that are Brønsted-Lowry acids?

Individual questions for students:

- Are all the reactions in Table 1 that are Brønsted-Lowry bases?

Group questions:

- Are all the reactions in Table 1 that are Brønsted-Lowry acids?

Individual questions for students:

- Are all the reactions in Table 1 that are Brønsted-Lowry bases?

Class:

- Are all the reactions in Table 1 that are Brønsted-Lowry acids?

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Clifford W. Hand, Harry Lyon Blewitt



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1. Understanding the eBook Chemquest 48 Intro To Acids And Bases
 - The Rise of Digital Reading Chemquest 48 Intro To Acids And Bases
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemquest 48 Intro To Acids And Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 48 Intro To Acids And Bases
 - Personalized Recommendations
 - Chemquest 48 Intro To Acids And Bases User Reviews and Ratings
 - Chemquest 48 Intro To Acids And Bases and Bestseller Lists
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- Chemquest 48 Intro To Acids And Bases Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chemquest 48 Intro To Acids And Bases Compatibility with Devices
 - Chemquest 48 Intro To Acids And Bases Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemquest 48 Intro To Acids And Bases
 - Highlighting and Note-Taking Chemquest 48 Intro To Acids And Bases
 - Interactive Elements Chemquest 48 Intro To Acids And Bases
- 8. Staying Engaged with Chemquest 48 Intro To Acids And Bases
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemquest 48 Intro To Acids And Bases
- 9. Balancing eBooks and Physical Books Chemquest 48 Intro To Acids And Bases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemquest 48 Intro To Acids And Bases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemquest 48 Intro To Acids And Bases
 - Setting Reading Goals Chemquest 48 Intro To Acids And Bases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemquest 48 Intro To Acids And Bases
 - Fact-Checking eBook Content of Chemquest 48 Intro To Acids And Bases
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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