

ANSWERS TO TOPIC 1

Review Questions

1. 1 2. 3 3. 2
4. 4 5. 2 6. 2
7. 1 8. 3 9. 3
10. The plum pudding model pictures the atom with an even distribution of positive charges, with electrons throughout the atom. The planetary model pictures the nucleus at the center of the atom with electrons in orbits.
11. Rutherford fired positively charged alpha particles at gold foil. He expected them to pass through with little deflection. When some were greatly deflected, he reasoned that they must have been repelled by a massive positively charged particle, which was called the nucleus.
12. The Rutherford model pictured the electrons in fixed, orbiting positions. The wave-mechanical model pictures electrons as possessing both wave and particle properties.
13. 3 14. 3 15. 1
16. 3 17. 4 18. 4
19. 2 20. 4 21. 2
22. 1 23. 1 24. 3
25. 3 26. 2 27. 2
28. 3 29. 3 30. 2
31. 3 32. 4 33. 4
34. 4 35. 1 36. 2
37. 1 38. 1 39. 3
40. 1 41. 2 42. 3
43. 2 44. 4 45. 4
46. 3 47. 3 48. 4
49. 2 50. 3 51. 2
52. 2 53. 3 54. 1
55. 3 56. 3
57. Na $1s^2 2s^2 2p^6 3s^1$
Mg $1s^2 2s^2 2p^6 3s^2$
Al $1s^2 2s^2 2p^6 3s^2 3p^1$
Si $1s^2 2s^2 2p^6 3s^2 3p^2$
P $1s^2 2s^2 2p^6 3s^2 3p^3$
S $1s^2 2s^2 2p^6 3s^2 3p^4$
Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
58. 1 59. 4 60. 1
61. 1 62. 2 63. 4
64. 4 65. 2 66. 2
67. 2 68. 2 69. 1
70. The mixture contains at least three parts. One solid remained on the filter paper. The other two parts consist of a liquid with a solid dissolved in it.
71. If the calcium and sulfur can be separated by physical means, the substance is a mixture. If chemical means are required to separate them, it is a compound.
72. The picture should show several of each type of molecule evenly mixed.
73. A mixture of iron and oxygen would have the iron present as a solid and the oxygen present as a gas. The iron would be attracted to a magnet. A glowing splint would ignite into a flame if placed into the oxygen gas. If the iron and oxygen were

combined as a compound, it would be a solid that is not attracted to a magnet.

74. Pepperoni pizza is heterogeneous because individual parts can be seen.
75. A is a mixture of an element and a compound; one molecule contains identical atoms and the other one contains different types of atoms. B is a compound; all particles are identical and are made up of two different types of atoms. C is an element; all molecules are identical and each is made up of two identical atoms. D is a mixture of an element and two different compounds; one molecule contains identical atoms and the other two contain different types of atoms, but they differ from each other.
76. Container D is homogeneous because the particles are mixed on the molecular level.
77. The two substances are different materials because their composition varies.

Regents Practice Questions

Part A

- | | | |
|-------|-------|-------|
| 1. 1 | 2. 1 | 3. 1 |
| 4. 3 | 5. 1 | 6. 4 |
| 7. 2 | 8. 3 | 9. 3 |
| 10. 2 | 11. 3 | 12. 1 |
| 13. 2 | 14. 3 | |

Part B

- | | | |
|-------|-------|-------|
| 15. 4 | 16. 3 | 17. 1 |
| 18. 3 | 19. 1 | 20. 3 |
| 21. 1 | 22. 3 | 23. 2 |
| 24. 4 | | |

Part C

25. Because the two nuclei have different numbers of protons, they have different amounts of attraction for electrons. As a result, the amounts of energy separating the energy levels differ for the two elements.
26. The samples are not the same substance because the percent of each element differs from sample to sample.
27. Because he assumed the atom to have an even distribution of positive charges, he did not expect a concentration of positive charges that would strongly deflect or bounce back the alpha particles.
28. The spectra are based on the elements present in the compound. If some lines matched and some didn't, the compounds have a least one element in common and at least one differing element.
29. Column 1: 1, 2, 1; column 2: K, 39, 19; column 3: 92, 146, 92
30. $(0.9778)(12) + (0.0222)(13) = 12.02$
31. They contain different numbers of neutrons.
32. $(19.99 \times .909) + (20.99 \times 0.03) + (21.99 \times 0.88)$
33. 20
34.
$$\frac{(12.02 \times 9.93) + (13.00 \times 1.07)}{100}$$
35. Both isotopes contain 6 protons but different numbers of neutrons. Or: An isotope with a mass of 12.02 contains 6 protons and 6 neutrons, while the isotope with a mass of 13.00 contains 6 protons and 7 neutrons.

Chemquest 15 Answer

S Marginson



Chemquest 15 Answer:

Making the Unseen Real Mark Vincent Falzon, 2009 **Virtual ADMET Assessment in Target Selection and Maturation** Bernard Testa, Lechoslaw Turski, 2006 A significant contributing factor to the evolution in drug discovery was the methodological and technological revolution with the advent of combinatorial chemistry This volume summarizes discussions of three aspects of modern drug discovery that is priority for targets early ADMET assessment and in silico screening **Data Base Alert** ,1986 Chemistry and Industry ,1990 Chemical Information Management Wendy A. Warr, Claus Suhr, 1992 *Directory of Online Databases* ,1991 Even numbered issues will contain update information odd numbered issues will be complete reissues with all new revised information fully integrated into the basic Directory

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Table of Contents Chemquest 15 Answer

1. Understanding the eBook Chemquest 15 Answer
 - The Rise of Digital Reading Chemquest 15 Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemquest 15 Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemquest 15 Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 15 Answer
 - Personalized Recommendations
 - Chemquest 15 Answer User Reviews and Ratings
 - Chemquest 15 Answer and Bestseller Lists

5. Accessing Chemquest 15 Answer Free and Paid eBooks
 - Chemquest 15 Answer Public Domain eBooks
 - Chemquest 15 Answer eBook Subscription Services
 - Chemquest 15 Answer Budget-Friendly Options
6. Navigating Chemquest 15 Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemquest 15 Answer Compatibility with Devices
 - Chemquest 15 Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemquest 15 Answer
 - Highlighting and Note-Taking Chemquest 15 Answer
 - Interactive Elements Chemquest 15 Answer
8. Staying Engaged with Chemquest 15 Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemquest 15 Answer
9. Balancing eBooks and Physical Books Chemquest 15 Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemquest 15 Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemquest 15 Answer
 - Setting Reading Goals Chemquest 15 Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemquest 15 Answer
 - Fact-Checking eBook Content of Chemquest 15 Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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