

# Gases and Moles

Date: \_\_\_\_\_

Hour: \_\_\_\_\_

## Information: The Gas Constant

We previously examined the fact that  $\frac{PV}{T}$  is a constant as long as the number of moles of a gas remains unchanged.

This means that any time you measure a gas, the pressure times the volume divided by the temperature (in Kelvin) will always equal the same number unless the amount of gas in the container changes. We are now going to explore how the pressure, volume and temperature depend on the number of moles of gas that are present.

Table 1: Experimental Data for Three Different Samples of Gas That Contain the Same Amount of Gas (# of moles of gas is constant)

Gas	# of moles of Gas in container (mol)	Volume of Container (L)	Temperature of Gas (K)	Pressure of Gas (kPa)
A	0.3008	14.0	672.0	120.0
B	0.3008	16.0	1280	200.0
C	0.3008	10.0	340.0	85.0

Table 2: Experimental Data for Three Different Samples of Gases That Contain Different Amounts of Gas (# of moles of gas is not constant)

Gas	# of moles of Gas in container	Volume of Container (L)	Temperature of Gas (K)	Pressure of Gas (kPa)
D	3.3485	40.0	230.0	160.0
E	0.7306	22.5	315	85
F	0.85950	16.0	280	125

## Critical Thinking Questions

1. Prove that  $\frac{PV}{T}$  is a constant for the data in Table 1, but not a constant for the data in Table 2.

2. Why is  $\frac{PV}{T}$  a constant in Table 1, but not in Table 2?

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Gases and Gas Laws Hobsons Publishing, PLC, 1983-01-01      **An Introduction to the Gas Phase** Claire

Vallance, 2017-11-23 An Introduction to the Gas Phase is adapted from a set of lecture notes for a core first year lecture course in physical chemistry taught at the University of Oxford The book is intended to give a relatively concise introduction to the gas phase at a level suitable for any undergraduate scientist After defining the gas phase properties of gases such as temperature pressure and volume are discussed The relationships between these properties are explained at a molecular level and simple models are introduced that allow the various gas laws to be derived from first principles Finally the collisional behaviour of gases is used to explain a number of gas phase phenomena such as effusion diffusion and thermal conductivity

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