

ChemQuest 48

pH and pOH

$$\text{pH} = \text{pOH} + \text{pK}_w$$

Information: Molecular Weights

There are two kinds of pH, one for acids and one for bases which will be very different when doing calculations. The first is based on the concentration of hydrogen ion (H^+) in solution. The other is one of hydroxide ions (OH^-) in solution and. They both use the same logarithm but the other (OH^-) is measured and reported.

The pH of a solution is $\text{pH} = -\log[\text{H}^+]$ and the pOH is $\text{pOH} = -\log[\text{OH}^-]$

Calculations: Examples

1. Calculate the pH of a solution of $1.5 \times 10^{-3} \text{ M}$ H^+ ions. $\text{pH} = -\log(1.5 \times 10^{-3}) = 2.82$
2. Calculate the pOH of a solution of $1.5 \times 10^{-3} \text{ M}$ OH^- ions. $\text{pOH} = -\log(1.5 \times 10^{-3}) = 2.82$
3. Calculate the pH of 0.1 M H^+ ions. $\text{pH} = -\log(0.1) = 1.00$

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Victor M. Corman



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