

# Intro. to Chemical Reactions

Name \_\_\_\_\_

Date \_\_\_\_\_

Class \_\_\_\_\_

## Formula Review

1. Write the formulas for the following compounds:

- a. aluminum phosphate \_\_\_\_\_ b. sodium phosphate \_\_\_\_\_  
 c. copper (II) phosphate \_\_\_\_\_ d. sodium sulfate \_\_\_\_\_

2. Write the formulas for the following compounds:

- a.  $\text{CaCl}_2$  \_\_\_\_\_ b.  $\text{FeCl}_3$  \_\_\_\_\_  
 c.  $\text{FeCl}_2$  \_\_\_\_\_ d.  $\text{CuCl}_2$  \_\_\_\_\_

## Chemical Reactions: Virtual Energy Unit 10a

Write down observations for the reaction as it occurs:



There is no observable gas or color change when these aqueous solutions:

Copper(II) ions react with hydroxide and silver ions precipitate and silver ions precipitate.

Write a balanced chemical equation for the reaction between hydroxide and silver ions. Include the states of matter.

## Chemical Reaction Equations

3. Write the following reaction under the correct state (gas, liquid, solid):



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**Yi-Tong Ma**

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