

# Ionic Bonding

Name: \_\_\_\_\_  
Hour: \_\_\_\_\_

## Information: Polyatomic Ions

The word, "polyatomic" means "many atoms". A polyatomic ion, therefore, is an ion that is made of more than one atom. An example of a polyatomic ion is the sulfate ion,  $\text{SO}_4^{2-}$ . Sulfate is composed of one sulfur atom and four oxygen atoms and overall sulfate has a negative two charge.

Some polyatomic ions:

Sulfate:  $\text{SO}_4^{2-}$ Cyanide:  $\text{CN}^-$ Acetate:  $\text{C}_2\text{H}_3\text{O}_2^-$ Phosphate:  $\text{PO}_4^{3-}$ Ammonium:  $\text{NH}_4^+$ Hydroxide:  $\text{OH}^-$ Nitrate:  $\text{NO}_3^-$ Chlorate:  $\text{ClO}_3^-$ Carbonate:  $\text{CO}_3^{2-}$ 

## Critical Thinking Questions

1. What do all of the polyatomic ions that have the suffix "-ate" have in common?
2. Which two atoms do you think compose the polyatomic ion called "silicate"?
3. What is the difference between calcium nitride and calcium nitrate?

## Information: Writing Formulas With Polyatomic Ions

First of all, you must remember that you can never change the formula for a polyatomic ion. Sulfate is always  $\text{SO}_4^{2-}$  and never  $\text{S}_2\text{O}_8^{4-}$  or something else. Following are some examples of chemical formulas that contain polyatomic ions.

Ammonium chloride is formed from one ammonium ion ( $\text{NH}_4^+$ ) and one chloride ion ( $\text{Cl}^-$ ) to give the formula:  $\text{NH}_4\text{Cl}$ . Sodium sulfate requires two sodium ions ( $\text{Na}^+$ ) because sulfate ( $\text{SO}_4^{2-}$ ) has a negative two charge; the formula is:  $\text{Na}_2\text{SO}_4$ .

Consider calcium hydroxide. Calcium has a positive two charge ( $\text{Ca}^{2+}$ ) and hydroxide has a negative one charge ( $\text{OH}^-$ ). We need two hydroxide ions to combine with one calcium ion so that the overall charge ends up being zero. We write calcium hydroxide like  $\text{Ca}(\text{OH})_2$ .

Following are some more examples:

potassium acetate:  $\text{KC}_2\text{H}_3\text{O}_2$ barium phosphate:  $\text{Ba}_3(\text{PO}_4)_2$ magnesium nitrate:  $\text{Mg}(\text{NO}_3)_2$ calcium carbonate:  $\text{CaCO}_3$

# Chemquest Ion Practice

**Corwin Hansch, Peter George  
Sammes, John Bodenhan Taylor**



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