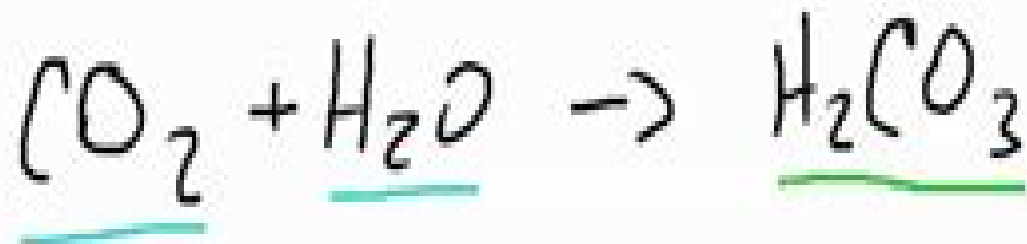


H_f : Heat of formation

$$\text{Sum } H_f \text{ Product - Sum } H_f \text{ reactants}$$



H_f will
be
given

$$H_f \text{ H}_2\text{CO}_3 = -275.2 \text{ KJ/mol}$$

$$H_f \text{ CO}_2 = -393.5 \text{ KJ/mol}$$

$$H_f \text{ H}_2\text{O} = -241.8 \text{ KJ/mol}$$

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AW Rasmussen



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