

Configuration Writing Practice

A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na: $1s^2, 2s^2, 2p^6, 3s^1$
2. Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{10}$
3. Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$
4. U: $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^{10}, 4p^6, 5s^2, 4d^{10}, 5p^6, 6s^2, 4f^{14}, 5d^{10}, 6p^6, 7s^2, 5f^6$
5. N: $1s^2, 2s^2, 2p^3$

B. Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example: N^{3-} is $1s^2 2s^2 2p^6$. It has three extra electrons

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11. O^{2-} : $1s^2, 2s^2, 2p^6$
12. Fe^{2+} : $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
13. B^{3+} : $1s^2$
14. Ni^{2+} : $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^8$
15. K^+ : $1s^2 2s^2 2p^6 3s^2 3p^6$
16. Co^{3+} : $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$

C. For the following electron configurations determine the possible elements (or ions) they may represent

17. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$: Se
18. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$: Kr
19. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{10}$: Dy
20. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$: Mn
21. [Kr] $5s^2 4d^{10} 5p^3$: As
22. [Ar] $4s^1$: K
23. [Xe] $6s^2 4f^{10}$: Dy
24. [Ne] $3s^2 3p^1$: Al

C. In the space below, write the full (unabbreviated) electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$
- 5) neptunium $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^{10}, 4p^6, 5s^2, 4d^{10}, 5p^6, 6s^2, 4f^{14}, 5d^{10}, 6p^6, 7s^2, 5f^6$

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