

## Pre-Lab for Determination of Solubility Product Constant

1. Write the solubility product constant ( $K_{sp}$ ) for the following reaction as a function of the concentrations of the products:



2.  $\text{Ca}(\text{IO}_3)_2$  will ionize in water to produce  $\text{IO}_3^-$  ions. The  $\text{IO}_3^-$  ions will react with KI. Write the reaction for this reaction.
3. We will be using starch as an indicator. Why?
4. In this experiment, we will first produce  $\text{I}_2$  by mixing calcium iodate with excess amount of KI. Based on the following reactions, how many moles of  $\text{I}_2$  will be produced if we started with 6 moles of calcium iodate and excess amount of KI?



5. In this experiment, we will titrate  $\text{I}_2$  with  $\text{Na}_2\text{S}_2\text{O}_3$  (sodium thiosulfate). The reaction is provided below. If we needed 22.5 mL of 0.15 M  $\text{Na}_2\text{S}_2\text{O}_3$  to reach the endpoint, how many moles of  $\text{I}_2$  were present in the solution? **(Must show all calculation steps.)**



6. How many moles of  $\text{Ca}(\text{IO}_3)_2$  are needed to produce the amount of  $\text{I}_2$  in Problem #5?
7. The solubility product constant ( $K_{sp}$ ) for calcium iodate is  $6.47 \times 10^{-6}$ . The molar mass of calcium iodate is 389.88 g/mol. 175.4 g of calcium iodate is added to 1 liter of water. What would be the concentrations of  $\text{Ca}^{2+}$  and  $\text{IO}_3^-$  at equilibrium? Would this solution form a precipitate? Is the solution unsaturated, saturated, or supersaturated when it reaches equilibrium? **(Must show all calculation steps.)**

# Determination Of A Solubility Product Lab Answers

**K Morrison**



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