

Determination of an Equilibrium Constant

Introduction

Chemical equilibrium is a state at which both the concentration of the product(s) and reactant(s) remain unchanged with time. In this state, the rate of forward reaction is equal to the rate of reverse reaction. This relationship can be quantified using the equilibrium constant, K_c as demonstrated:



This expression goes to show that in order to reach equilibrium, each species in the reaction must satisfy the equilibrium constant. If it is not satisfied, it can be deduced that the reaction is still ongoing until it will reach the point of equilibrium. The purpose of this experimentation is for students to determine K_c using the stoichiometry of a reaction.

Procedure

For this experimentation, students observed the reaction between Fe^{3+} and SCN^- to find the equilibrium constant of the reaction:



This was demonstrated in a Youtube video in order for students to learn how it would be done in a laboratory. Students made five test tubes with varying concentrations of the reactants, $\text{Fe}(\text{NO}_3)_3$, KSCN , and diluted with HClO_4 , as shown in Table 1. An additional sixth test tube was made and used as a standard. The solutions were mixed well in test tubes and set aside. Students were to make sure not to cross contaminate by washing and drying the stirring rod before use.

The purpose of the HClO_4 was to prevent the formation of a red precipitate $\text{Fe}(\text{OH})_3$. Ideally we wanted the reactants to be clear so that what was observed and measured was FeSCN^{2+} which is a blood-red colored species. Because it is colored, its concentration can be measured using a spectrometer, in which the instrument measures its light absorbance. After the spectrometer was calibrated, the instrument was set to absorbance mode then students began to measure the absorbance of each test tube. All results and calculations were noted.

Determination Of Equilibrium Constant Lab Report

Conclusion

Yi-Tong Ma



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