

Date Performed: 9/18/06  
Date Submitted: 9/25/06  
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## Determination of a Chemical Formula

### Objective:

To determine the percentage composition and simplest formula for an oxide of copper by its reduction with methane gas ( $\text{CH}_4$ ) at approximately  $500^\circ\text{C}$ . This experiment is a study of stoichiometry.

### Experimental Data:

Mass of Test Tube	$17.726 \pm .001 \text{ g}$
Mass of Test Tube and $\text{Cu}_2\text{O}$	$19.705 \pm .001 \text{ g}$
Mass of $\text{Cu}_2\text{O}$	$1.979 \pm .001 \text{ g}$
Mass of Test Tube and Cu	$19.313 \pm .001 \text{ g}$
Mass of Cu	$1.587 \pm .001 \text{ g}$
Mass of O	$0.392 \pm .001 \text{ g}$

### Sample Calculation:

Mass percentage calculation:

$$\frac{\text{Mass of Atom}}{\text{Mass of Compound}} \times 100 = \frac{1.587 \pm .001 \text{ g of Cu}}{1.979 \pm .001 \text{ g of Cu}_2\text{O}} \times 100 = 79.73\%$$

Calculation of a Chemical Formula:

Assume a 100 g compound and calculate individual mass using mass percentage.

$$79.73 \text{ g of Cu} \times \frac{1 \text{ mol of Cu}}{63.55 \text{ g per mol}} = 1.255 \text{ mol of Cu}$$

$$20.07 \text{ g of O} \times \frac{1 \text{ mol of O}}{16.00 \text{ g per mol}} = 1.254 \text{ mol of O}$$

### Results and Conclusions:

Mass of Test Tube	$17.726 \pm .001 \text{ g}$
Mass of Test Tube and $\text{Cu}_2\text{O}$	$19.705 \pm .001 \text{ g}$
Mass of $\text{Cu}_2\text{O}$	$1.979 \pm .001 \text{ g}$
Mass of Test Tube and Cu	$19.313 \pm .001 \text{ g}$
Mass of Cu	$1.587 \pm .001 \text{ g}$
Mass of O	$0.392 \pm .001 \text{ g}$
Mass Percentage of Cu	79.73%
Mass Percentage of O	20.07%

Cu and O have a one to one mole ratio; therefore the empirical formula is  $\text{CuO}$ .

# Determining A Chemical Formula Lab Report

**R Sanford**



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## **Table of Contents Determining A Chemical Formula Lab Report**

1. Understanding the eBook Determining A Chemical Formula Lab Report
  - The Rise of Digital Reading Determining A Chemical Formula Lab Report
  - Advantages of eBooks Over Traditional Books
2. Identifying Determining A Chemical Formula Lab Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Determining A Chemical Formula Lab Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from Determining A Chemical Formula Lab Report
  - Personalized Recommendations
  - Determining A Chemical Formula Lab Report User Reviews and Ratings
  - Determining A Chemical Formula Lab Report and Bestseller Lists
5. Accessing Determining A Chemical Formula Lab Report Free and Paid eBooks
  - Determining A Chemical Formula Lab Report Public Domain eBooks
  - Determining A Chemical Formula Lab Report eBook Subscription Services
  - Determining A Chemical Formula Lab Report Budget-Friendly Options
6. Navigating Determining A Chemical Formula Lab Report eBook Formats
  - ePub, PDF, MOBI, and More
  - Determining A Chemical Formula Lab Report Compatibility with Devices
  - Determining A Chemical Formula Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Determining A Chemical Formula Lab Report
  - Highlighting and Note-Taking Determining A Chemical Formula Lab Report
  - Interactive Elements Determining A Chemical Formula Lab Report

8. Staying Engaged with Determining A Chemical Formula Lab Report
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Determining A Chemical Formula Lab Report
9. Balancing eBooks and Physical Books Determining A Chemical Formula Lab Report
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Determining A Chemical Formula Lab Report
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Determining A Chemical Formula Lab Report
  - Setting Reading Goals Determining A Chemical Formula Lab Report
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Determining A Chemical Formula Lab Report
  - Fact-Checking eBook Content of Determining A Chemical Formula Lab Report
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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