

Determining Empirical Formulas

1) 75% carbon, 25% hydrogen	Answer:	CH_4
Work: $\text{C } 75\% \rightarrow 75\text{g C} \left \frac{1 \text{ mol}}{12.01 \text{ g C}} = 6.24 \text{ mol} / 6.24 = 1 \right.$ $\text{H } 25\% \rightarrow 25\text{g H} \left \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 24.75 \text{ mol} / 6.24 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;">CH_4</div>		
2) 52.7% potassium, 47.3% chlorine	Answer:	KCl
Work: $\text{K } 52.7\% \rightarrow 52.7\text{g K} \left \frac{1 \text{ mol K}}{39.10 \text{ g K}} = 1.34 \text{ mol} / 1.33 = 1 \right.$ $\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left \frac{1 \text{ mol Cl}}{35.45 \text{ g Cl}} = 1.33 \text{ mol} / 1.33 = 1 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;">KCl</div>		
3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen	Answer:	AlPO_4
Work: $\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = 0.819 \text{ mol} / 0.819 = 1 \right.$ $\text{P } 25.4\% \rightarrow 25.4\text{g P} \left \frac{1 \text{ mol P}}{30.97 \text{ g P}} = 0.820 \text{ mol} / 0.819 = 1 \right.$ $\text{O } 52.5\% \rightarrow 52.5\text{g O} \left \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 3.28 \text{ mol} / 0.819 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;">AlPO_4</div>		
4) 13% magnesium, 87% bromine	Answer:	MgBr_2
Work: $13\text{g Mg} \left \frac{1 \text{ mol Mg}}{24.31 \text{ g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1 \right.$ $87\text{g Br} \left \frac{1 \text{ mol Br}}{79.90 \text{ g Br}} = 1.09 \text{ mol Br} / 0.54 = 2 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;">MgBr_2</div>		
5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen	Answer:	Na_2SO_4
Work: $32.4\text{g Na} \left \frac{1 \text{ mol}}{22.99 \text{ g Na}} = 1.41 \text{ mol Na} / 0.70 = 2 \right.$ $22.5\text{g S} \left \frac{1 \text{ mol S}}{32.37 \text{ g S}} = 0.70 \text{ mol S} / 0.70 = 1 \right.$ $45.1\text{g O} \left \frac{1 \text{ mol O}}{16 \text{ g O}} = 2.82 \text{ mol O} / 0.70 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;">Na_2SO_4</div>		

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Table of Contents Determining Empirical Formulas Page 30 Answers

1. Understanding the eBook Determining Empirical Formulas Page 30 Answers
 - The Rise of Digital Reading Determining Empirical Formulas Page 30 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Determining Empirical Formulas Page 30 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Determining Empirical Formulas Page 30 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Determining Empirical Formulas Page 30 Answers
 - Personalized Recommendations
 - Determining Empirical Formulas Page 30 Answers User Reviews and Ratings
 - Determining Empirical Formulas Page 30 Answers and Bestseller Lists
5. Accessing Determining Empirical Formulas Page 30 Answers Free and Paid eBooks

- Determining Empirical Formulas Page 30 Answers Public Domain eBooks
 - Determining Empirical Formulas Page 30 Answers eBook Subscription Services
 - Determining Empirical Formulas Page 30 Answers Budget-Friendly Options
6. Navigating Determining Empirical Formulas Page 30 Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Determining Empirical Formulas Page 30 Answers Compatibility with Devices
 - Determining Empirical Formulas Page 30 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Determining Empirical Formulas Page 30 Answers
 - Highlighting and Note-Taking Determining Empirical Formulas Page 30 Answers
 - Interactive Elements Determining Empirical Formulas Page 30 Answers
8. Staying Engaged with Determining Empirical Formulas Page 30 Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Determining Empirical Formulas Page 30 Answers
9. Balancing eBooks and Physical Books Determining Empirical Formulas Page 30 Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Determining Empirical Formulas Page 30 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Determining Empirical Formulas Page 30 Answers
- Setting Reading Goals Determining Empirical Formulas Page 30 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Determining Empirical Formulas Page 30 Answers
- Fact-Checking eBook Content of Determining Empirical Formulas Page 30 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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