



## EXPERIMENT 1: DETERMINING MOLAR VOLUME

## Data Sheet

Table 3: Amount of  $\text{H}_2\text{O}$  Displaced for Each Trial-Grams of Solute Used

Trial	Barometric Pressure (atm)	Temperature ( $^{\circ}\text{C}$ )	Mass of Chemical (g)	Water Displaced (mL)
Trial 1 $\text{NaHCO}_3$	.99	19.6	1.0	83
Trial 2 $\text{NaHCO}_3$	.99	20.2	.9	68
Trial 3 $\text{Na}_2\text{CO}_3$	.99	20.0	1.3	96
Trial 4 $\text{Na}_2\text{CO}_3$	.99	20.0	1.0	105

## Post-Lab Questions

1. What is the average experimental volume per mole of carbon dioxide calculated for sodium carbonate and sodium bicarbonate?
2. List two sources of error in your experiment. What did they affect? Explain how they could have been fixed.
3. Identify two assumptions that must be true when using the ideal gas equation that are not true for real gases.
4. Why is the air in a hot air balloon heated?

# Determining The Molar Volume Of A Gas Lab Worksheet

**Thomas Engel**



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KEY BENEFIT Physical Chemistry for the Life Sciences presents the core concepts of physical chemistry with mathematical rigor and conceptual clarity and develops the modern biological applications alongside the physical principles The traditional presentations of physical chemistry are augmented with material that makes these chemical ideas biologically relevant applying physical principles to the understanding of the complex problems of 21st century biology  
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