

## EXPERIMENT 2

## Determination of a Chemical Formula

## Additional Review Material

- Relevant sections in the text (*Chemistry for Engineering Students*, 2<sup>nd</sup> Ed., Brown, Holme): 3.2 – 3.5
- Techniques (see Blackboard/Laboratory/Techniques): Vacuum Filtration (You Tube)

## Background

The law of definite proportions states that a compound always contains two or more elements combined in the same definite proportion by mass. The chemical formula that represents this relative number of atoms in a compound (the simplest mole ratio) is called the empirical formula. The chemical formula that gives the actual number of atoms in a compound is called the molecular formula. Consider for example, a compound composed of two atoms, H and O. We know that water consists of two atoms of hydrogen and a single oxygen. Its molecular formula is  $H_2O$ . Another compound composed of these same elements, but in different proportions, is hydrogen peroxide. The molecular formula for hydrogen peroxide is  $H_2O_2$  and the empirical formula, specifying the smallest whole-number ratio of atoms in the compound, is  $HO$ .

In the laboratory, the empirical formula of a compound can be determined by measuring the mass of each component in a compound. In this experiment, the empirical formula of a solid compound with the general formula  $Cu_xCl_y \cdot zH_2O$  will be determined by measuring the mass of  $H_2O$ , copper (Cu), and chlorine (Cl) in the compound. The values  $x$ ,  $y$ , and  $z$  represent integers which will establish the proper chemical formula.

This type of ionic compound is called a hydrate. Hydrates have a specific number of water molecules associated with each formula unit. When writing these chemical formulas, the formula of the anhydride is followed by a raised dot which represents the "weak" bond between the anhydride and the number of water molecules (also called water of hydration). Hydrates and the corresponding anhydrous compounds often have distinctly different physical and chemical properties and their ability to gain or lose water makes them versatile and useful in a variety of practical applications.

To determine the percent composition of water in the unknown  $Cu_xCl_y \cdot zH_2O$  compound, water of hydration is first removed by heating the sample.

$$\%H_2O \text{ in compound} = \left( \frac{\text{mass of } H_2O}{\text{original mass of compound}} \right) (100)$$

Remaining  $Cu_xCl_y$  salts then readily dissolve to form aqueous solutions of the constituent ions:



We can now use the oxidation-reduction (redox) reaction between copper ions and solid magnesium (Mg) to convert  $Cu^{2+}$  to  $Cu(s)$ :



# Determination Of A Chemical Formula Experiment Answers

**Michal Rosen-Zvi**



## Determination Of A Chemical Formula Experiment Answers:

*Chemistry in the Laboratory* James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations **Solutions, Phase**

**equilibrium, Conductance, Electrochemistry and Functional Group Organic Chemistry II - Laboratory** Mr. Rohit Manglik, 2024-03-04 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels

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Pharmaceutical Journal , 1900 **Transactions of the Pharmaceutical Meetings** , 1889 **A Text-book of Ordnance and Gunnery** Lawrence Laurensen Bruff, 1896 *Appletons' Annual Cyclopedia and Register of Important Events* , 1901 **The Chemical News and Journal of Physical Science** , 1903 **Report of the Superintendent of**

**Public Instruction of the State of New York** New York (State). Department of Public Instruction, 1902 **The Systematic Identification of Organic Compounds** Christine K. F. Hermann, Terence C. Morrill, Ralph L. Shriner, Reynold C. Fuson, 2023-03-08 The Systematic Identification of Organic Compounds A comprehensive introduction to the identification of unknown organic compounds Identifying unknown compounds is one of the most important parts of the study of chemistry From basic characteristics such as melting and or boiling point to more complex data generated through cutting edge techniques the range of possible methods for identifying unknown organic compounds is substantial The utility of a research reference which compiles known techniques and characteristics of possible compounds is clear The Systematic Identification of Organic Compounds provides such a reference designed to teach a hands on approach in the chemistry lab It takes readers step by step through the process of identifying an unknown compound and elucidating its structure from infrared nuclear magnetic resonance and mass spectra in addition to solubility characteristics melting point boiling point and classification tests The result is an essential overview for advanced chemistry students looking to understand this exciting area of laboratory work Readers of the ninth edition of The Systematic Identification of Organic Compounds will also find A detailed chapter on safety personal protection equipment chemical storage safety data sheets and other safety concerns New NMR IR and mass spectra with detailed explanations on interpretation Questions at the end of each chapter designed to facilitate and reinforce progression keyed to a companion website for instructors Tables of known compounds including data relevant for identification Companion website with structural problems from experimental data for students to practice how

to reason and solve The Systematic Identification of Organic Compounds is a useful reference for advanced undergraduates and graduate students studying organic chemistry organic spectroscopy and related subjects     *The Pharmaceutical Journal* ... ,1889     **Chemical Principles** Peter Atkins,Loretta Jones,2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry s frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding     *The pharmaceutical journal and transactions* ,1889     The Software Encyclopedia 2000 Bowker Editorial Staff,2000-05     **Appletons' Annual Cyclopaedia and Register of Important Events** ,1901     **The Chemical News and Journal of Industrial Science** William Crookes,James H. Gardiner,Gerald Druce,H. W. Blood-Ryan,1898     Appletons' Annual Cyclopædia and Register of Important Events ... ,1901     Journal of the Chemical Society ,1905     Journal of the Chemical Society Chemical Society (Great Britain),1905

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