

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Solve the problems below:

1. The empirical formula of a compound is PCl_2 . Its molecular mass is 263 g/mol. What is its molecular formula?

Solution given by student is:

Answer

2. The empirical formula of a compound is CH_2 . Its molecular mass is 78 g/mol. What is its molecular formula?

Solution given by student is:

Answer

3. A compound is found to contain 69.8% carbon, 4.9% hydrogen and 25.3% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Answer

4. A compound is found to contain 69.8% carbon, 4.9% hydrogen and 25.3% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Answer

5. A compound is found to contain 69.8% carbon, 4.9% hydrogen and 25.3% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Answer

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