

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Solve the problems below:

1. The empirical formula of a compound is PCl_2 . Its molecular mass is 263 g/mol. What is its molecular formula?

Solution given by student is:

2. The empirical formula of a compound is CH_2 . Its molecular mass is 78 g/mol. What is its molecular formula?

Solution given by student is:

3. A compound is found to contain 69.8% carbon, 6.7% hydrogen and 23.5% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

4. A compound is found to contain 69.8% carbon, 6.7% hydrogen and 23.5% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

5. A compound is found to contain 69.8% carbon, 6.7% hydrogen and 23.5% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Determining Molecular Formulas True Formulas Answer Key

Ralph Thomas Myers



Determining Molecular Formulas True Formulas Answer Key:

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